

PCT

WORLD INTELLECTUAL PROPERTY ORGANIZATION  
International Bureau



3. 11. 1994  
PCT/IE 800/0663

INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

<b>(51) International Patent Classification <sup>5</sup> :</b> C07F 7/18, 7/12, 7/21 C08G 77/50, 77/52 C09D 183/14	<b>A1</b>	<b>(11) International Publication Number:</b> WO 94/06807 <b>(43) International Publication Date:</b> 31 March 1994 (31.03.94)
<b>(21) International Application Number:</b> PCT/US93/08685 <b>(22) International Filing Date:</b> 15 September 1993 (15.09.93)  <b>(30) Priority data:</b> 07/945,777 16 September 1992 (16.09.92) US 08/120,995 13 September 1993 (13.09.93) US  <b>(71) Applicant:</b> E.I. DU PONT DE NEMOURS AND COMPANY [US/US]; 1007 Market Street, Wilmington, DE 19898 (US).  <b>(72) Inventors:</b> MICHALCZYK, Michael, Joseph ; 16 Servan Court, Wilmington, DE 19805 (US). SHARP, Kenneth, George ; 164 Hamilton Road, Landenburg, PA 19350-9357 (US).		<b>(74) Agents:</b> MAYER, Nancy, S. et al.; E.I. du Pont de Nemours and Company, Legal/Patent Records Center, 1007 Market Street, Wilmington, DE 19898 (US).  <b>(81) Designated States:</b> CA, JP, European patent (AT, BE, CH, DE, DK, ES, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE).  <b>Published</b> <i>With international search report.</i>
<b>(54) Title:</b> SINGLE COMPONENT INORGANIC/ORGANIC NETWORK MATERIALS AND PRECURSORS THERE-OF  <b>(57) Abstract</b>  Single component inorganic/organic network materials incorporating the physical properties of glasses with the flexibility of organic materials of empirical formula $X(\text{SiO}_{1.5})_n$ , wherein X is one or more flexible organic linkages and n is greater than or equal to 2, as well as precursors thereof, are disclosed.		

*FOR THE PURPOSES OF INFORMATION ONLY*

Codes used to identify States party to the PCT on the front pages of pamphlets publishing international applications under the PCT.

AT	Austria	FR	France	MR	Mauritania
AU	Australia	GA	Gabon	MW	Malawi
BB	Barbados	GB	United Kingdom	NE	Niger
BE	Belgium	GN	Guinea	NL	Netherlands
BF	Burkina Faso	GR	Greece	NO	Norway
BG	Bulgaria	HU	Hungary	NZ	New Zealand
BJ	Benin	IE	Ireland	PL	Poland
BR	Brazil	IT	Italy	PT	Portugal
BY	Belarus	JP	Japan	RO	Romania
CA	Canada	KP	Democratic People's Republic of Korea	RU	Russian Federation
CF	Central African Republic	KR	Republic of Korea	SD	Sudan
CG	Congo	KZ	Kazakhstan	SE	Sweden
CH	Switzerland	LI	Liechtenstein	SI	Slovenia
CI	Côte d'Ivoire	LK	Sri Lanka	SK	Slovak Republic
CM	Cameroon	LU	Luxembourg	SN	Senegal
CN	China	LV	Latvia	TD	Chad
CS	Czechoslovakia	MC	Monaco	TC	Togo
CZ	Czech Republic	MG	Madagascar	UA	Ukraine
DE	Germany	ML	Mali	US	United States of America
DK	Denmark	MN	Mongolia	UZ	Uzbekistan
ES	Spain			VN	Viet Nam
FI	Finland				

TITLESINGLE COMPONENT INORGANIC/ORGANIC NETWORK  
MATERIALS AND PRECURSORS THEREOF

This invention concerns single component  
5 inorganic/organic materials which consist essentially of  
a multicomponent network comprising flexible organic and  
rigid inorganic portions.

A significant limitation on both the generation and  
utility of inorganic network materials such as glasses  
10 and ceramics relates to their brittleness. When glasses  
are prepared at room temperatures using sol-gel  
technology, Brinker, C. J., et al., Sol Gel Science,  
Academic Press, San Diego, CA (1990) drying stresses  
cause catastrophic fracture of films more than about  
15 0.5 micron in thickness (for fully dense silica); larger  
monolithic structures are possible only with low drying  
rates. Applications for inorganic glasses are limited  
to those in which considerable amounts of energy  
absorption or dissipation (i.e., toughness) are not  
20 required.

Prior attempts to make the glass network more  
compliant have involved limiting the number of  
networking bonds per silicon atom (e.g., using  
alkyl(trialkoxysilanes instead of tetraalkoxy silanes).  
25 There is growing interest in inorganic/organic hybrid  
materials which incorporate both glasses and flexible  
organic material. One approach has been to incorporate  
organic polymers into silica glasses.

K. J. Shea et al., Chemistry of Materials, 1,  
30 572(1989), disclose organically modified silicates  
prepared by sol-gel processing of bis-triethoxysilylaryl  
and bis-trichlorosilylaryl monomers. The three monomers  
employed had the aryl portion of the monomer as  
phenylene (-C<sub>6</sub>H<sub>4</sub>-), biphenylene (-C<sub>6</sub>H<sub>4</sub>-C<sub>6</sub>H<sub>4</sub>-), and  
35 triphenylene (-C<sub>6</sub>H<sub>4</sub>-C<sub>6</sub>H<sub>4</sub>-C<sub>6</sub>H<sub>4</sub>-). Rigid networks are

produced which are brittle, porous and contain a single rigid organic link between silicon atoms.

The present invention comprises a new class of network materials which incorporate both glasses and flexible organic materials without suffering the deficiencies of the glasses produced by conventional sol-gel technology. This invention concerns certain compositions having chemically bonded inorganic network portions and organic network portions. These two portions may be derived from a single precursor molecule or from a mixture of precursor molecules, which precursor molecules contain the elements of, or precursors to the elements of, both the inorganic and organic portions. Because both the organic and inorganic portions of the composition derive from the same precursor molecule, or from a mixture of such precursor molecules, the portions cannot be separated without the breaking of chemical bonds.

#### SUMMARY OF THE INVENTION

This invention comprises an inorganic/organic composition of the idealized empirical formula (II):



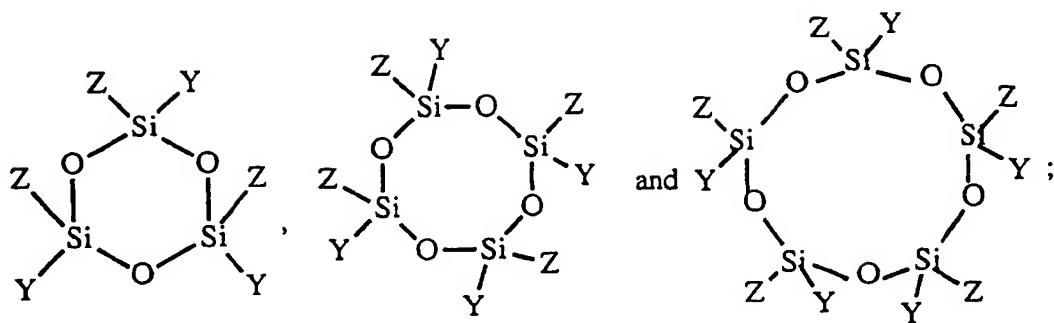
wherein

n is an integer greater than or equal to 2; and  
X is at least one flexible organic link selected

from the group consisting of:

- (a)  $\text{R}^1_m\text{SiY}_{4-m}$ ;
- (b) ring structures

3



IIa

IIb

IIc

- (c)  $R^1_m Si (OSi (CH_3)_2 Y)_{4-m}$ ;  
 (d)  $R^1_m Si (OY)_{4-m}$ ;  
 (e)  $CH_3 Si Y_2 - O - Si Y_2 CH_3$ ;  
 (f)  $Y (CH_3)_2 Si - C_6H_4 - Si (CH_3)_2 Y$ ;  
 (g)  $O [-C_6H_4 - Si (CH_3)_2 Y]_2$ ;  
 (h)  $O [Si (CH_3)_2 Y]_2$ ;  
 (i)  $Y (CH_3)_2 Si CH_2 - CH_2 Si (CH_3)_2 Y$ ;  
 (j)  $Y (CF_2)_p Y$ , provided that when p is 6, Y is other than ethylene;  
 (k)  $Y_3 Si O Si Y_3$ ;  
 (l)  $Y_3 Si (CH_2)_b Si Y_3$ ; and  
 (m)  $Y_3 Si C_6H_4 Si Y_3$ ;  
 (n) substituted benzene, including all isomers, selected from the group consisting of:  
 (i)  $C_6H_3 (SiZ_{3-a} Y_a)_3$ ;  
 (ii)  $C_6H_2 (SiZ_{3-a} Y_a)_4$ ;  
 (iii)  $C_6H (SiZ_{3-a} Y_a)_5$ ; and  
 (iv)  $C_6 (SiZ_{3-a} Y_a)_6$ ; and  
 (o) substituted cyclohexane, including all stereoisomers, selected from the group consisting of:  
 (i) 1,2- $C_6H_{10} (Y)_2$ ; 1,3- $C_6H_{10} (Y)_2$ ;  
 1,4- $C_6H_{10} (Y)_2$ ;  
 (ii) 1,2,4- $C_6H_9 (Y)_3$ ; 1,2,3- $C_6H_9 (Y)_3$ ;  
 1,3,5- $C_6H_9 (Y)_3$ ;

(iii) 1,2,3,4-C<sub>6</sub>H<sub>8</sub>(Y)<sub>4</sub>; 1,2,4,5-C<sub>6</sub>H<sub>8</sub>(Y)<sub>4</sub>;

1,2,3,5-C<sub>6</sub>H<sub>8</sub>(Y)<sub>4</sub>;

(iv) 1,2,3,4,5-C<sub>6</sub>H<sub>7</sub>(Y)<sub>5</sub>; and

(v) C<sub>6</sub>H<sub>6</sub>(Y)<sub>6</sub>;

5

wherein:

Z is an alkyl group of 1 to 4 carbon atoms,

3,3,3-trifluoropropyl, aralkyl, or aryl;

Y is (CR<sup>2</sup>R<sup>3</sup>)<sub>k</sub>CR<sup>4</sup>R<sup>5</sup>CR<sup>6</sup>R<sup>7</sup>(CR<sup>8</sup>R<sup>9</sup>)<sub>h</sub>-;

R<sup>1</sup> is alkyl of 1 to about 8 carbon atoms or

10

aryl;

R<sup>2</sup> to R<sup>9</sup> are each independently hydrogen, alkyl

of 1 to about 8 carbon atoms or aryl,

provided that at least one of R<sup>4</sup> to R<sup>7</sup> is

hydrogen;

15

m is 0, 1 or 2;

k and h are each independently an integer from

0 to 10, provided that at least one of k or

h is zero;

a is 1, 2 or 3;

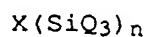
20

p is an even integer from 4 to 10; and

b is an integer from 1 to 10.

This invention also comprises:

a compound of the formula (I):



(I)

wherein:

25

Q is alkoxy of 1 to about 8 carbon atoms, acyloxy

of 1 to about 8 carbon atoms, or halogen;

n is an integer greater than or equal to 2; and

X is at least one flexible organic link selected

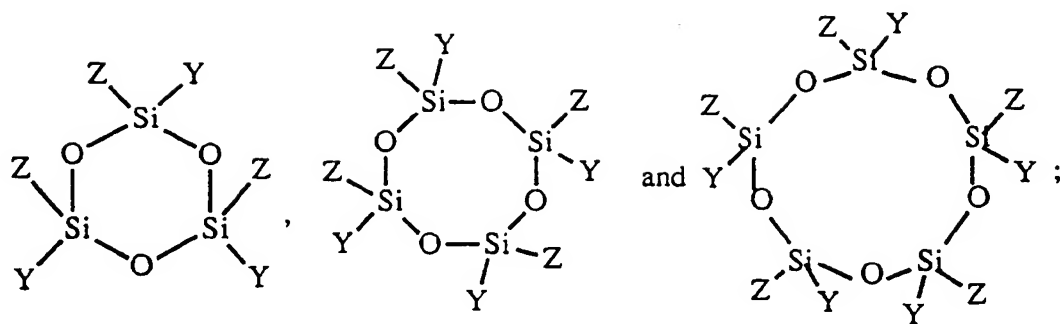
from the group consisting of:

30

(a) R<sup>1</sup><sub>m</sub>SiY<sub>4-m</sub>;

(b) ring structures

5

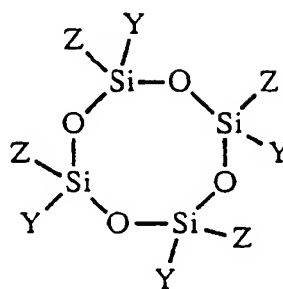


1a

1b

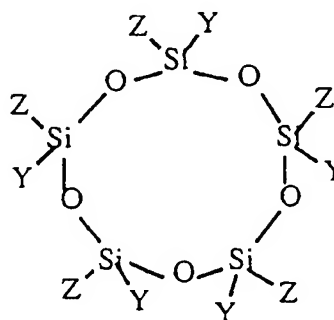
1c

provided that when X is



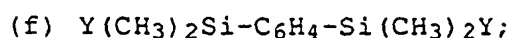
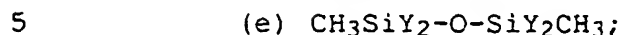
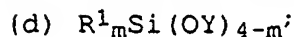
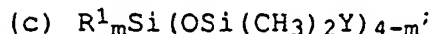
1b

Z is other than methyl and Y is other than ethylene or propylene; and when X is

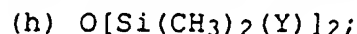


1c

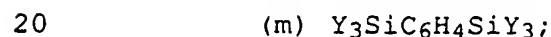
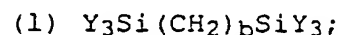
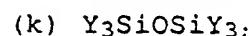
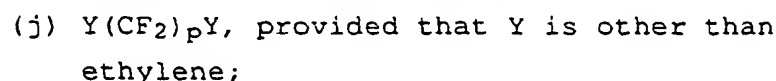
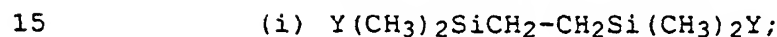
Z is other than methyl and Y is other than ethylene or propylene;



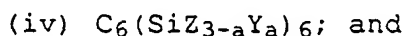
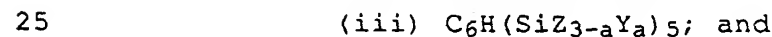
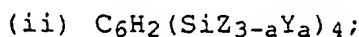
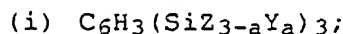
provided that in the definition of Y as defined below either h or k is greater than zero when Q is ethoxy;



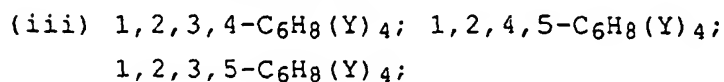
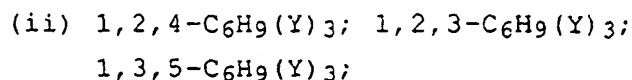
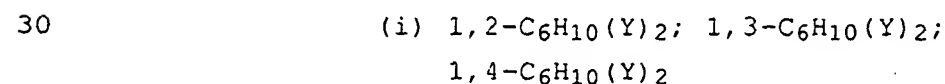
provided that in the definition of Y as defined below either h or k is greater than zero when Q is ethoxy;



(n) substituted benzene, including all isomers, selected from the group consisting of:



(o) substituted cyclohexane, including all stereoisomers, selected from the group consisting of:





(iv) 1,2,3,4,5-C<sub>6</sub>H<sub>7</sub>(Y)<sub>5</sub>; and

(v) C<sub>6</sub>H<sub>6</sub>(Y)<sub>6</sub>;

wherein:

Z is an alkyl group of 1 to 4 carbon atoms,  
3,3,3-trifluoropropyl, aralkyl or aryl;

Y is (CR<sup>2</sup>R<sup>3</sup>)<sub>k</sub>CR<sup>4</sup>R<sup>5</sup>CR<sup>6</sup>R<sup>7</sup>(CR<sup>8</sup>R<sup>9</sup>)<sub>h</sub>-;

R<sup>1</sup> is alkyl of 1 to about 8 carbon atoms or  
aryl;

R<sup>2</sup> to R<sup>9</sup> are each independently hydrogen, alkyl  
of 1 to about 8 carbon atoms or aryl,  
provided that at least one of R<sup>4</sup> to R<sup>7</sup> is  
hydrogen;

m is 0, 1 or 2;

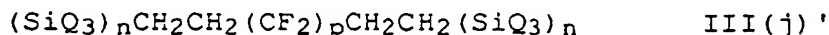
k and h are each independently an integer from  
0 to 10, provided that at least one of k or  
h is zero;

a is 1, 2 or 3;

p is an even integer from 4 to 10; and

b is an integer from 1 to 10.

This invention further comprises a compound of the  
formula III(j)':



wherein:

Q is alkoxy of 1 to about 8 carbon atoms,  
acyloxy of 1 to about 8 carbon atoms, or  
halogen;

n is an integer greater than or equal to 2;

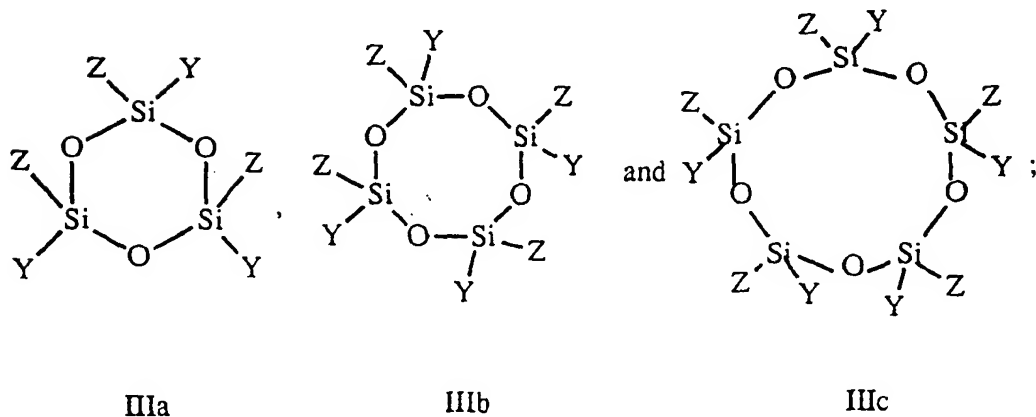
and p is an even integer from 4 to 10.

This invention further comprises a process for the  
preparation of a compound of formula (I), X(SiQ<sub>3</sub>)<sub>n</sub>, as  
defined above comprising reacting a compound containing  
an Si-H group with a compound containing an olefinic or

alkynyl bond in the presence of a transition metal catalyst such as platinum, or a free-radical initiator.

This invention further comprises a method for modifying sol-gel glasses comprising:

- 5 (a) combining a star gel precursor compound of formula (III),  $X(\text{SiQ}_3)_n$  wherein
  - Q is alkoxy of 1 to about 8 carbon atoms, acyloxy of 1 to about 8 carbon atoms, or halogen;
  - 10 n is an integer greater than or equal to 2; and
  - X is at least one flexible organic link selected from the group consisting of:
    - (a)  $\text{R}^1_m\text{SiY}_{4-m}$ ;
    - 15 (b) ring structures



- (c)  $\text{R}^1_m\text{Si}(\text{OSi}(\text{CH}_3)_2\text{Y})_{4-m}$ ;
- (d)  $\text{R}^1_m\text{Si}(\text{OY})_{4-m}$ ;
- (e)  $\text{CH}_3\text{SiY}_2\text{-O-SiY}_2\text{CH}_3$ ;
- (f)  $\text{Y}(\text{CH}_3)_2\text{Si-C}_6\text{H}_4\text{-Si}(\text{CH}_3)_2\text{Y}$ ;
- 20 (g)  $\text{O}[-\text{C}_6\text{H}_4\text{-Si}(\text{CH}_3)_2\text{Y}]_2$ ;
- (h)  $\text{O}[\text{Si}(\text{CH}_3)_2\text{Y}]_2$ ;
- (i)  $\text{Y}(\text{CH}_3)_2\text{SiCH}_2\text{-CH}_2\text{Si}(\text{CH}_3)_2\text{Y}$ ;
- (j)  $\text{Y}(\text{CF}_2)_p\text{Y}$ ;
- (k)  $\text{Y}_3\text{SiOSiY}_3$ ;

- (l)  $Y_3Si(CH_2)_bSiY_3$ ; and  
(m)  $Y_3SiC_6H_4SiY_3$ ;  
(n) substituted benzene, including all isomers, selected from the group consisting of:  
5           (i)  $C_6H_3(SiZ_{3-a}Y_a)_3$ ;  
            (ii)  $C_6H_2(SiZ_{3-a}Y_a)_4$ ;  
            (iii)  $C_6H(SiZ_{3-a}Y_a)_5$ ; and  
            (iv)  $C_6(SiZ_{3-a}Y_a)_6$ ; and  
10           (o) substituted cyclohexane, including all stereoisomers, selected from the group consisting of:  
            (i) 1,2- $C_6H_{10}(Y)_2$ ; 1,3- $C_6H_{10}(Y)_2$ ;  
                1,4- $C_6H_{10}(Y)_2$   
15           (ii) 1,2,4- $C_6H_9(Y)_3$ ; 1,2,3- $C_6H_9(Y)_3$ ;  
                1,3,5- $C_6H_9(Y)_3$ ;  
            (iii) 1,2,3,4- $C_6H_8(Y)_4$ ;  
                1,2,4,5- $C_6H_8(Y)_4$ ;  
                1,2,3,5- $C_6H_8(Y)_4$ ;  
20           (iv) 1,2,3,4,5- $C_6H_7(Y)_5$ ; and  
            (v)  $C_6H_6(Y)_6$ ;  
wherein:  
Z is an alkyl group of 1 to 4 carbon atoms,  
3,3,3-trifluoropropyl, aralkyl, or  
25           aryl;  
Y is  $(CR^2R^3)_kCR^4R^5CR^6R^7(CR^8R^9)_h-$ ;  
 $R^1$  is alkyl of 1 to about 8 carbon atoms or  
aryl;  
 $R^2$  to  $R^9$  are each independently hydrogen,  
30           alkyl of 1 to about 8 carbon atoms or  
aryl, provided that at least one of  $R^4$   
to  $R^7$  is hydrogen;  
m is 0, 1 or 2;

k and h are each independently an integer from 0 to 10, provided that at least one of k or h is zero;

a is 1, 2 or 3;

5 p is an even integer from 4 to 10; and

b is an integer from 1 to 10;

with a metal alkoxide sol-gel precursor;

(b) mixing in water with a solvent and a catalyst or a carboxylic acid optionally in the presence of a solvent; and

10 (c) drying.

The resulting modified sol-gel glass can tolerate increased drying rates and shows lower brittleness compared to the corresponding unmodified sol-gel glass.

15 This invention further comprises a process for the preparation of the composition of formula (II) as defined above comprising:

(a) mixing at least one compound of formula (I) or formula (III) as defined above with water in the presence of a solvent and a catalyst, or with at least one strong carboxylic acid having a pKa value of a maximum of about 4.0 and containing from 0 to 20 mole % water;

(b) maintaining the mixture resulting from step (a) at a temperature within the range of about 0-100°C; and

25 (c) isolating the resulting inorganic/organic composition of formula (II).

This invention further comprises a method for coating a substrate comprising reacting the star gel precursor of formula (III), as defined above, with water in the presence of a solvent and a catalyst, or a strong carboxylic acid optionally in the presence of a solvent, dipping the substrate in the resulting mixture, removing the coated substrate from the mixture, and drying the

coating to generate a substrate coated with a composition of formula II, as defined above.

DETAILED DESCRIPTION OF THE INVENTION

This invention provides certain inorganic/organic compositions comprising inorganic network portions and organic network portions. These two portions are derived from a single precursor molecule, or from a mixture of precursor molecules, which precursor molecules contain the elements of, or precursors to the elements of, both the inorganic and organic portions. Because both the organic and inorganic portions of the composition derive from the same precursor molecule, or from a mixture of such precursor molecules, the composition cannot be separated without the breaking of chemical bonds.

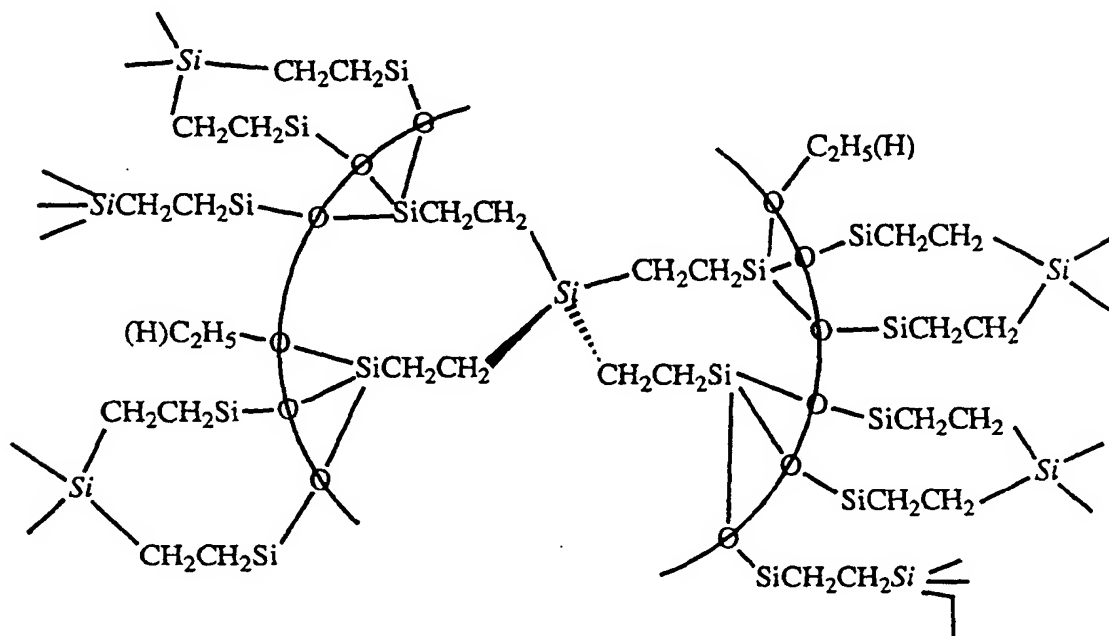
The present invention further comprises a method for modifying conventional sol-gel glasses to increase drying rates and lower brittleness comprising combining a star gel precursor of the present invention of formula (I) or formula (III) with a conventional sol-gel system based on tetraalkoxysilanes or other metal alkoxides; mixing in water with a solvent and a catalyst, or a carboxylic acid optionally in the presence of a solvent; and drying.

The inorganic/organic network compositions of the present invention, which can be in the form of gels or glasses, are of the idealized empirical formula (II):



as defined above wherein X is one or more flexible organic links, which simultaneously interconnect n silicon atoms where n is an integer greater than or equal to 2. Each of the latter atoms will be constituents of a network structure via bonds to other

silicon atoms through oxygen. For example, an inorganic/organic gel formed from Star 1 (Claim 1, (a)) which is  $X(\text{SiO}_{1.5})_n$  wherein  $X = \text{Si}(\text{CH}_2\text{CH}_2-)_4$ ,  $Y = -\text{CH}_2\text{CH}_2-$ ,  $m = 0$  and  $n = 4$  and the Si's of the Si-O-Si crosslinks are shown in standard print and the Si's of X in italics, could be represented as follows:



Condensation to form Si-O-Si cross links from Si-OC<sub>2</sub>H<sub>5</sub> by hydrolysis does not go to 100% completion; this is shown above by residual, uncrosslinked SiOC<sub>2</sub>H<sub>5</sub> or SiOH groups. An idealized formula corresponds to 100% crosslinking. The idealized empirical formula of the inorganic/organic gel derived from Star 1 as shown above would be:  $\text{Si}(\text{CH}_2\text{CH}_2\text{SiO}_{1.5})_4$  - the unit within the two half circles which bisect the oxygen atoms.

The number of SiO<sub>1.5</sub> groups depend on the number of Y's as defined in formula (I), (II) or (III). On the average there are 1.5 oxygens associated with each Si. For example, when there are two Y's there are 2 Si's.

Every Y is attached to a Si; there are no unsatisfied valences. There will be -OR or -OH groups that are not crosslinked which is desirable on the perimeter of the network for reacting with components in other

5 compositions.

These compositions are prepared by the hydrolysis of one or more star gel precursors of the present invention of formula (I) or formula (III). Star gel precursors are molecules which comprise a flexible  
10 organic or inorganic core comprising a central atom, ring or short linear segment linked to multiple arms which terminate in a silicon atom which bears at least two hydrolyzable substituents. The star gel precursors of the present invention comprise compounds of  
15 formula (I)

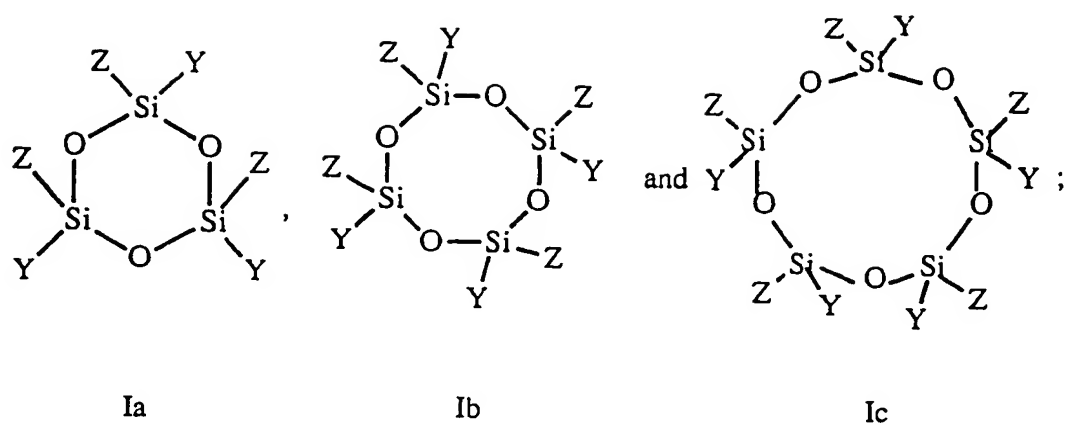


wherein X is at least one flexible organic link, as defined below, n is an integer greater than or equal to 2, and Q is a hydrolyzable group such as alkoxy containing from 1 to about 8 carbon atoms, acyloxy of 1  
20 to about 8 carbon atoms, or halogen.

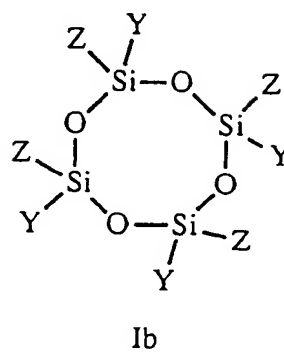
In general for formula (I), X comprises a central atom, ring or short linear segment with a number of arms which terminate in a silicon atom. In particular X comprises one or more flexible organic links selected  
25 from the group consisting of:

- (a)  $\text{R}^1_m\text{SiY}_{4-m}$ ;
- (b) ring structures

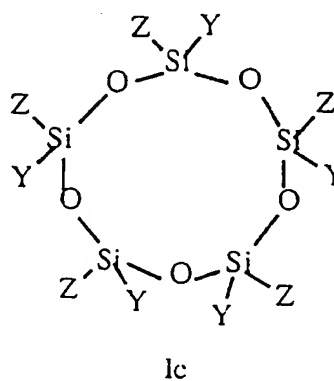
14



provided that when X is

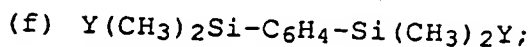
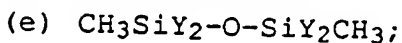
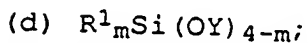
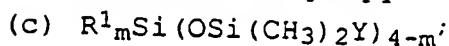


Z is other than methyl and Y is other than ethylene or propylene; and when X is

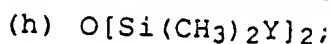
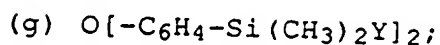




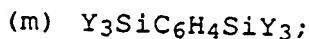
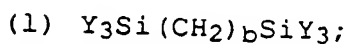
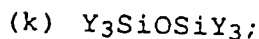
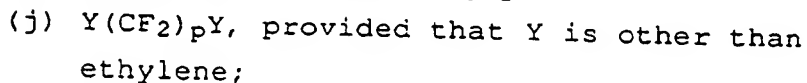
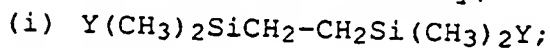
Z is other than methyl and Y is other than ethylene or propylene;



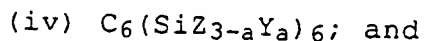
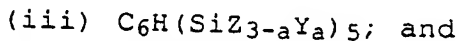
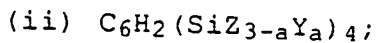
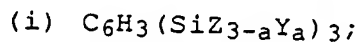
provided that in the definition of Y as defined below either h or k is greater than zero when Q is ethoxy;



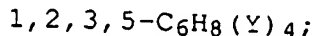
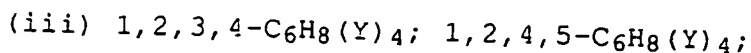
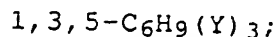
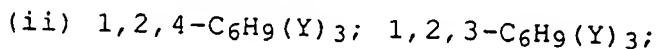
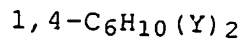
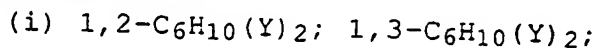
provided that in the definition of Y as defined below either h or k is greater than zero when Q is ethoxy;



(n) substituted benzene, including all isomers selected from the group consisting of:



(o) substituted cyclohexane, including all stereoisomers, selected from the group consisting of:



(iv) 1,2,3,4,5-C<sub>6</sub>H<sub>7</sub>(Y)<sub>5</sub>; and

(v) C<sub>6</sub>H<sub>6</sub>(Y)<sub>6</sub>;

wherein:

Z is an alkyl group of 1 to 4 carbon atoms,  
3,3,3-trifluoropropyl, aralkyl or aryl;

Y is (CR<sup>2</sup>R<sup>3</sup>)<sub>k</sub>CR<sup>4</sup>R<sup>5</sup>CR<sup>6</sup>R<sup>7</sup>(CR<sup>8</sup>R<sup>9</sup>)<sub>n</sub>-;

R<sup>1</sup> is alkyl of 1 to about 8 carbon atoms or  
aryl;

R<sup>2</sup> to R<sup>9</sup> are each independently hydrogen, alkyl  
of 1 to about 8 carbon atoms or aryl,  
provided that at least one of R<sup>4</sup> to R<sup>7</sup> is  
hydrogen;

m is 0, 1 or 2;

k and h are each independently an integer from

0 to 10, provided that at least one of k or  
h is zero;

a is 1, 2 or 3;

p is an even integer from 4 to 10; and

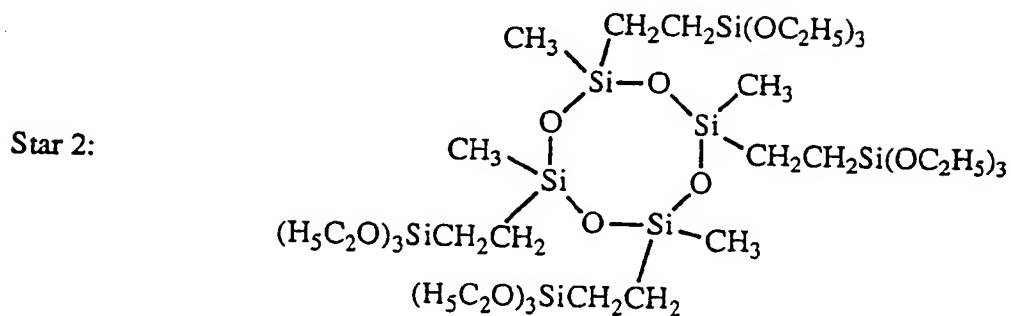
b is an integer from 1 to 10.

For formula (I), (II) and (III) the most preferred  
flexible organic link, X, is where m is 0, k is 0 or 1,  
h is 0 or 1, and all of R<sup>2</sup> to R<sup>9</sup> are hydrogen. The  
preferred Q are alkoxy of 1 to about 3 carbon atoms.  
Most preferred Q is ethoxy. The most preferred halogen  
is chloro. The preferred aralkyl is benzyl. The  
preferred aryl is phenyl.

Preferred star gel precursors of formula III  
include those listed in Table I below.

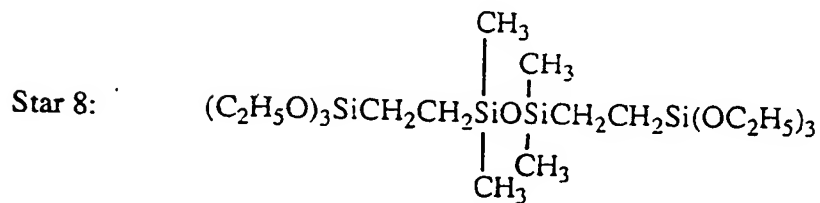
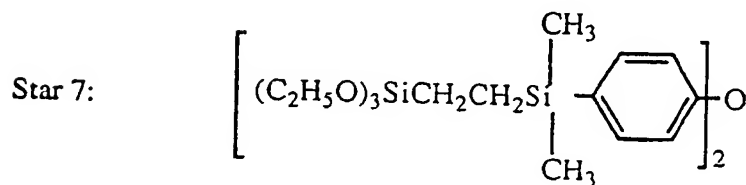
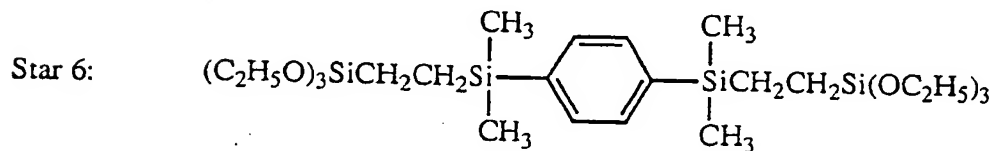
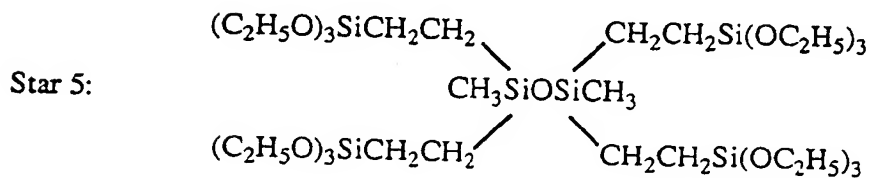
TABLE I  
Star-Gel Precursors

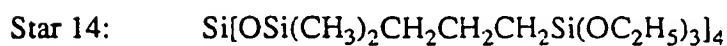
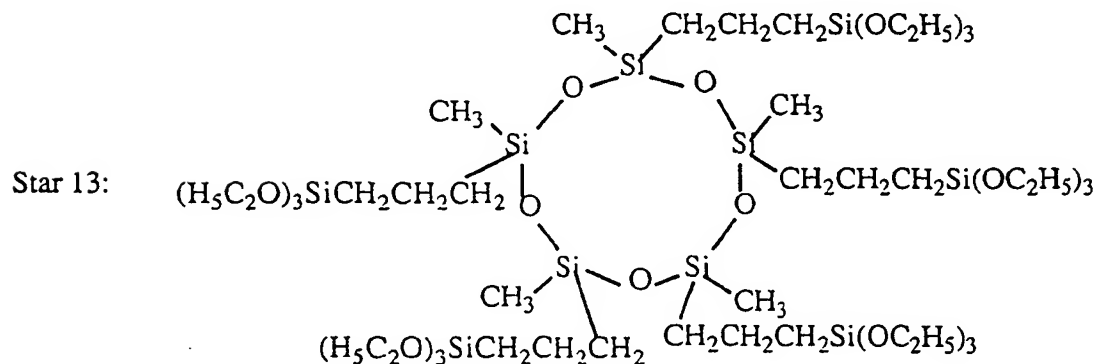
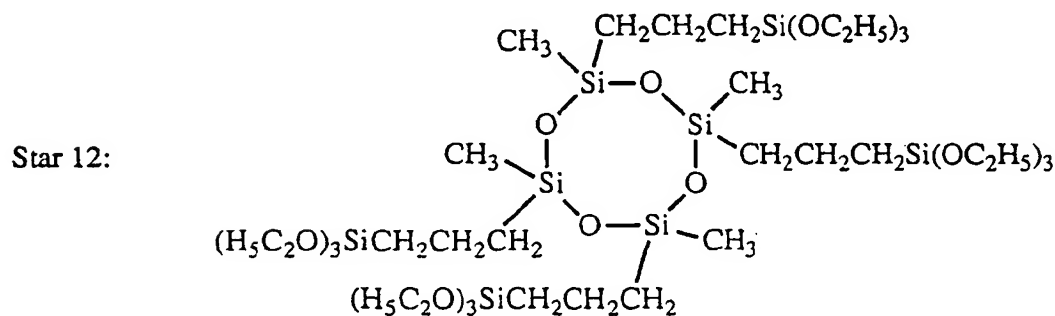
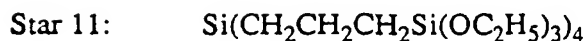
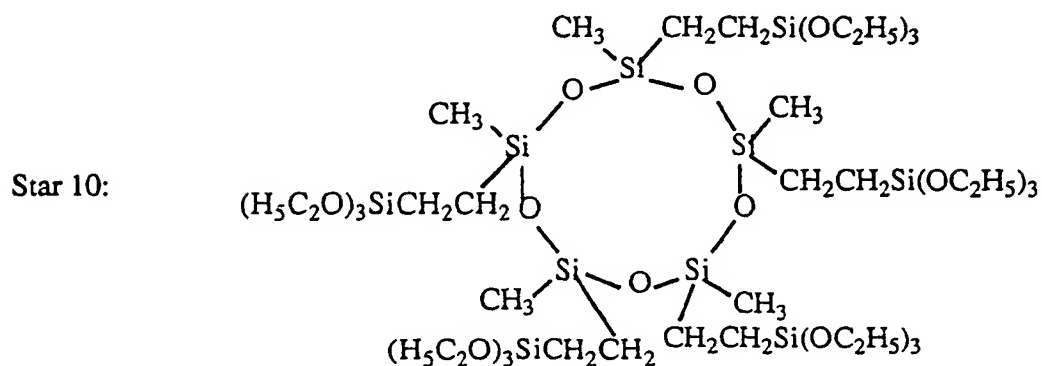
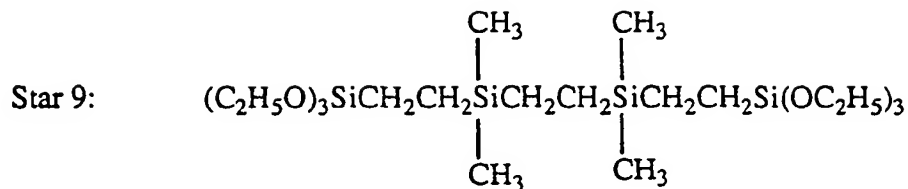
Star 1:  $\text{Si}(\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3)_4$

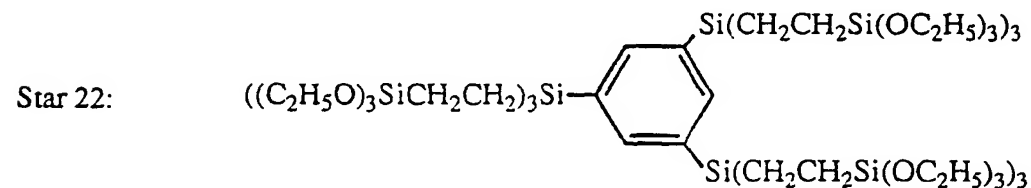
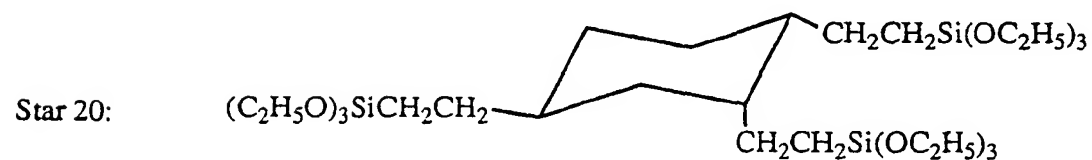
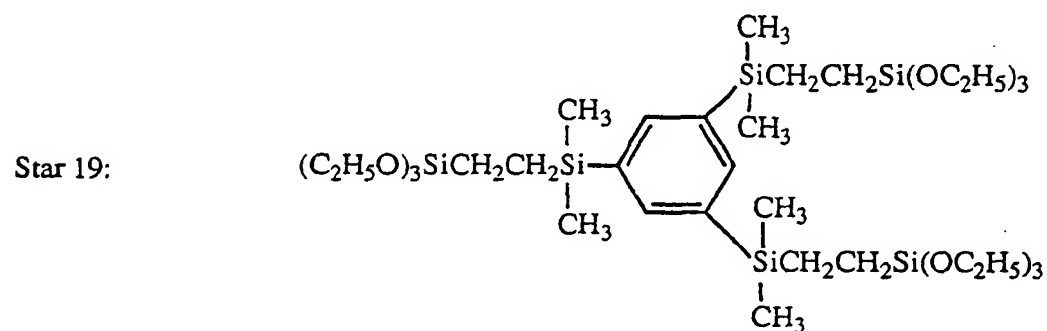
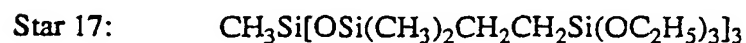
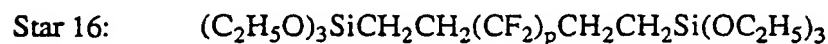
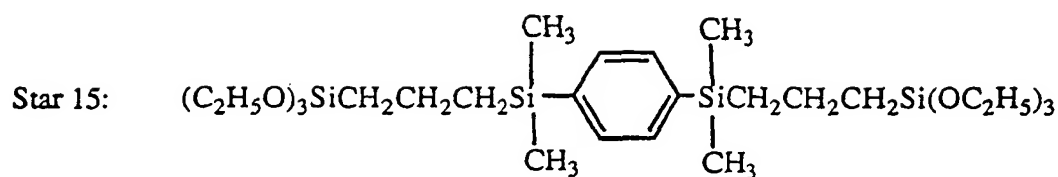


Star 3:  $\text{Si}[\text{OSi}(\text{CH}_3)_2\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3]_4$

Star 4:  $\text{Si}(\text{OCH}_2\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3)_4$

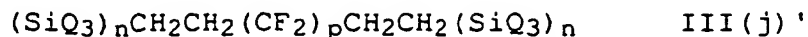






Preferred star gel precursors of formula I include Stars 1, 3, 4, 5, 7, 9, 11, 13, 14, 15, 17, 18, 19, 20, 21, and 22 as shown in Table I.

The present invention also comprises a compound of the formula III(j)':



5

wherein:

Q is alkoxy of 1 to about 8 carbon atoms,  
acyloxy of 1 to about 8 carbon atoms, or  
halogen;

10 n is an integer greater than or equal to 2; and  
p is an even integer from 4 to 10.

Q is preferably ethoxy or Cl. p is preferably 6 or 10.

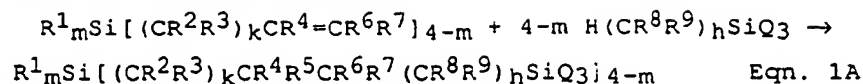
The present invention further comprises processes for preparation of star gel precursors of formula I as  
15 defined above.

Synthesis of the star gel precursors is afforded from hydrosilylation reactions, i.e. an addition reaction between a compound containing a Si-H group with a compound containing aliphatic unsaturation (C=C or  
20 -C≡C-) in the presence of a catalyst or free radical initiator. Precursor segments containing -CH=CH<sub>2</sub> groups react with other precursor segments which contain terminal Si-H bonds. With these precursor segments a number of different flexible star gel precursors can be  
25 constructed as illustrated in Table 1.

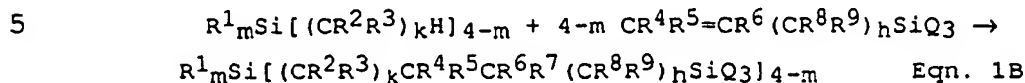
Either precursor segment may contain the vinyl or other unsaturated group capable of Si-H addition. For example, Si(CH=CH<sub>2</sub>)<sub>4</sub> reacts with HSi(OC<sub>2</sub>H<sub>5</sub>)<sub>3</sub> to form star gel precursor 1, Si[CH<sub>2</sub>CH<sub>2</sub>Si(OC<sub>2</sub>H<sub>5</sub>)<sub>3</sub>]<sub>4</sub>; and  
30 cyclo[(CH<sub>3</sub>)HSiO]<sub>5</sub> reacts with CH<sub>2</sub>=CH-Si(OC<sub>2</sub>H<sub>5</sub>)<sub>3</sub> to form star gel precursor 10, cyclo[OSi(CH<sub>3</sub>)CH<sub>2</sub>CH<sub>2</sub>Si(OC<sub>2</sub>H<sub>5</sub>)<sub>3</sub>]<sub>5</sub>.

All of the following equations with the exception of Equations 7B and 7C provide for preparation of compounds of formula I by addition of a silane across a  
35 carbon-carbon double bond for various definitions of X:

(a) when X is  $R^1_mSi[Y]_{4-m}$ :

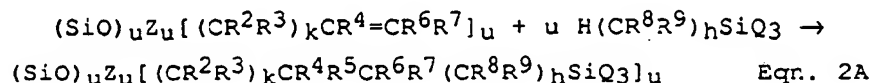


or

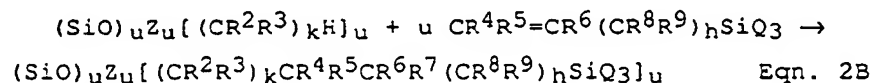


(b) when X is a ring structure of the type Ia, Ib or Ic as previously defined which can be abbreviated

10  $(SiO)_uZ_u(YSiQ_3)_u$ , wherein  $u = 3$  for Ia,  $u = 4$  for Ib, and  $u = 5$  for Ic; then

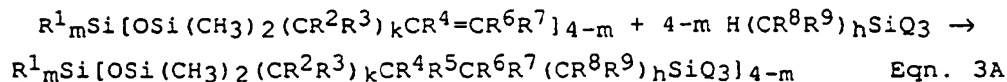


15

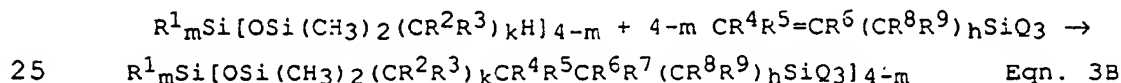


(c) when X is  $R^1_mSi[OSi(CH_3)_2Y]_{4-m}$ :

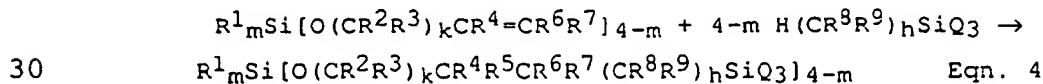
20



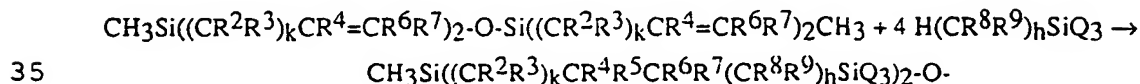
or



(d) when X is  $R^1_mSi[OY]_{4-m}$ :



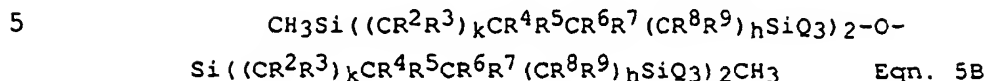
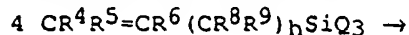
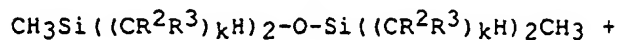
(e) when X is  $CH_3SiY_2-O-SiY_2CH_3$ :



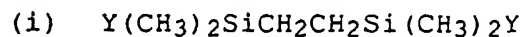
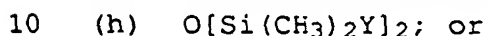
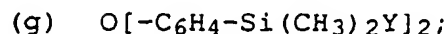
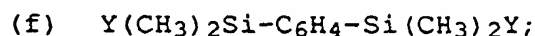
22



or

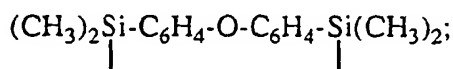
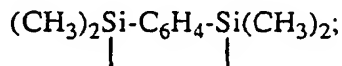


when X is

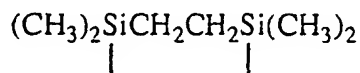
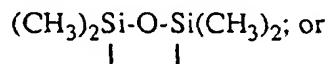


Formula I can generally be written as  $\text{D}(\text{YSiQ}_3)_2$  wherein Y is as previously defined and D is a connecting group chosen from:

15

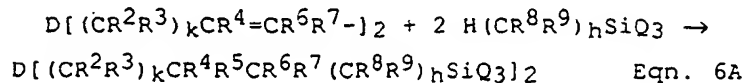


20

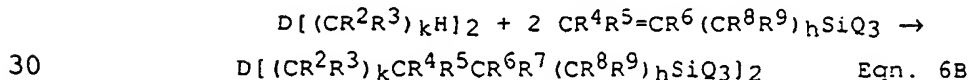


then the product  $\text{D}(\text{YSiQ}_3)_2$  is formed by the reaction

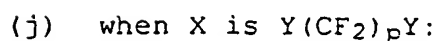
25



or

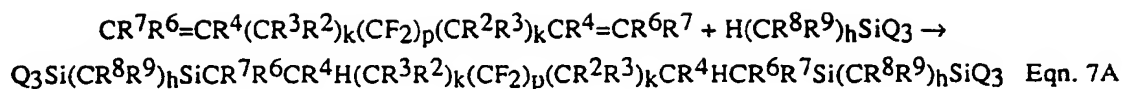


30

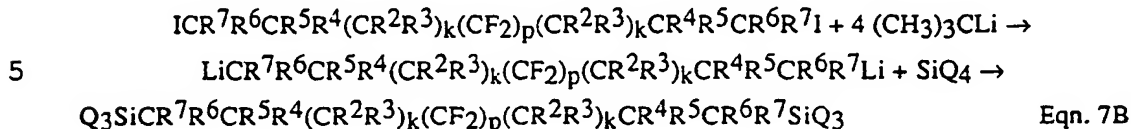




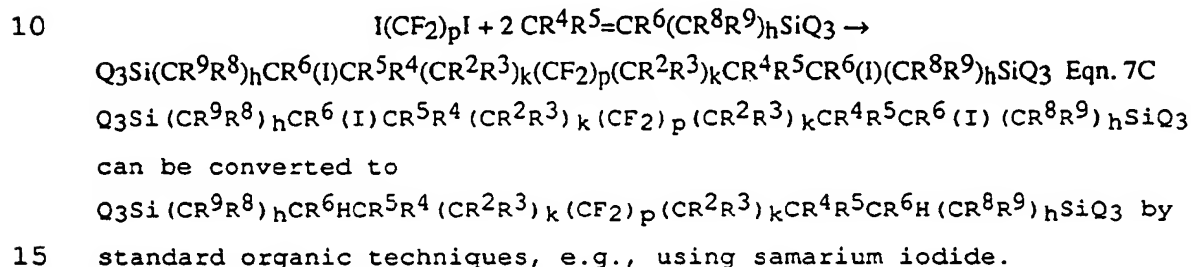
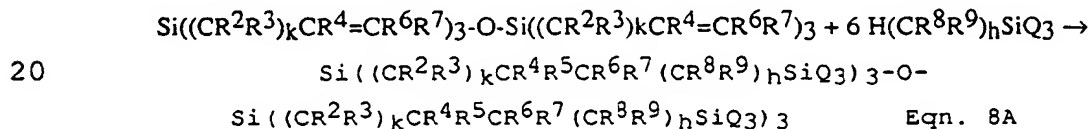
23



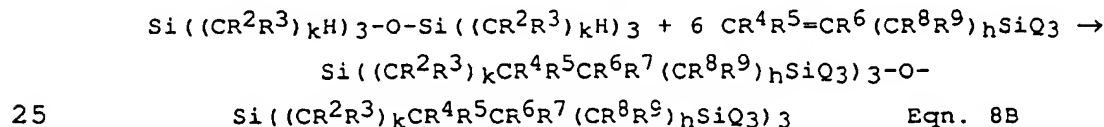
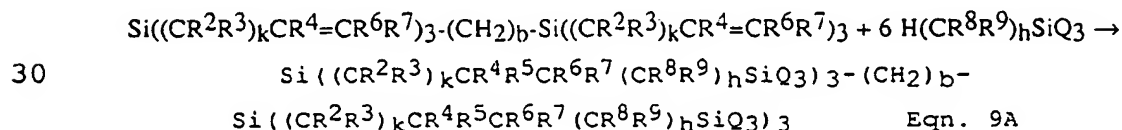
or



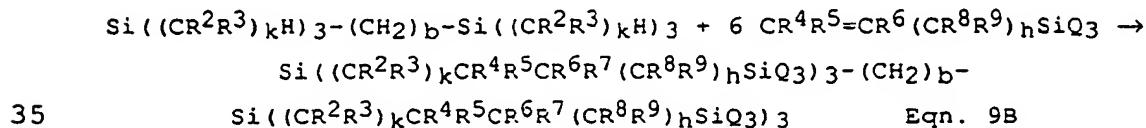
or

for  $k = 0$ ;  $p = 4, 6$  or  $8$ ; all R's = H:(k) when X is  $\text{Y}_3\text{-Si-O-Si-Y}_3$ :

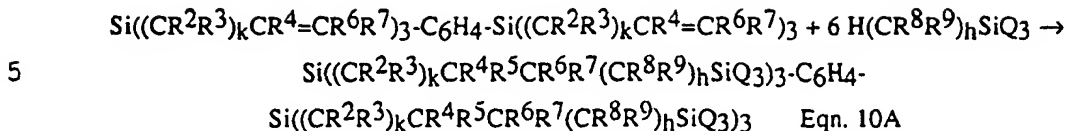
or

(l) when X is  $\text{Y}_3\text{-Si-(CH}_2)_b\text{-Si-Y}_3$ :

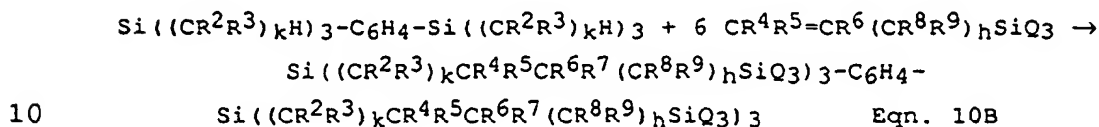
or



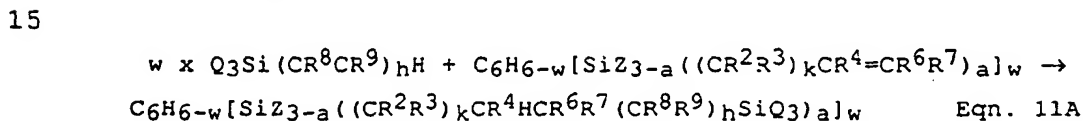
(m) when X is  $Y_3-Si-C_6H_4-Si-Y_3$ :



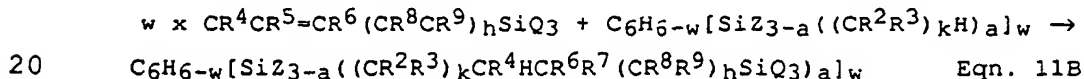
or



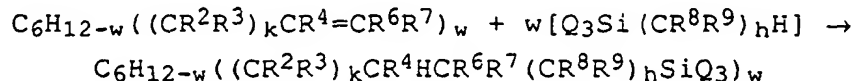
(n) when X is a substituted benzene structure of the type, as previously defined, which can be abbreviated  $C_6H_{6-w}(SiZ_{3-a}Y_a)_w$ :



or



(o) when X is a substituted cyclohexane structure of the type, as previously defined, which can be abbreviated  $C_6H_{12-w}(Y)_w$ , wherein w is the number of substituents; then



30 For convenience the reaction of Equations denoted A or B above is chosen depending upon the commercial availability of the starting reagents. In each set of equations where an A and B are presented, h = 0 in Eqn. A and k = 0 in Eqn. B. In (d), k ≥ 1 and h = 0.

35 Specific sources of reactants are listed hereinafter

just prior to the Examples. The reactants are employed in a ratio such that the precursor containing the  $\text{SiQ}_3$  group is employed in a molar excess of 10-50% to ensure completion of the hydrosilylation reaction. A

5 transition metal catalyst such as platinum, or a free radical initiator is employed in an effective amount. Examples of suitable free radical initiators include VAZO® compounds available from E. I. du Pont de Nemours and Company, Wilmington, DE.

10 These reactions can be conducted at a temperature of from about 25°C to about 100°C. Preferably the process is conducted at about 80°C to about 100°C. The pressure employed is typically ambient, about 1 atm ( $1.01 \times 10^5$  Pa). The reactions are carried out under an  
15 inert gas atmosphere, although use of an air atmosphere is not precluded. Reaction time is typically from about 4 hours to about 24 hours.

Use of solvent is not required in these reactions. Suitable solvents which may be employed are those  
20 capable of dissolving the reactants and which do not interfere with the reaction or generate unnecessary by-products. The desired product can be isolated by any means known to those skilled in the art. Preferably the desired product is isolated by removal of volatiles  
25 under reduced pressure.

NMR and  $\text{K}^+\text{IDS}$  mass spectrometry have been used to characterize product purities. Typically, yields of completely reacted material exceed 85%, with the principal impurities being either reverse (Markovnikov)  
30 hydrosilylation or incompletely substituted material containing unreacted  $-\text{CH}=\text{CH}_2$  groups. The catalyst can be removed, by filtering through silica gel or activated charcoal.

Synthesis of the star gel precursors wherein  
35  $\text{X} = \text{Y}(\text{CF}_2)_p\text{Y}$  may also be afforded from a metallation

reaction between an alpha-omega diiodoalkylperfluoro-alkane, e.g., contacted with tert-butyl lithium, followed by reaction with  $\text{Si}(\text{OEt})_4$ , as shown above in Equation 7B. Alternatively unsaturated trialkoxy-silanes, or trihalosilanes can be inserted into the C-I bond of  $\text{I}(\text{CF}_2)_p\text{I}$ , followed by reduction of the C-I to C-H using standard organic reduction reagents as shown in Equation 7C. Examples of suitable reagents are zinc metal, tri-n-butyl tin hydride or samarium iodide.

10        Possible uses of star gel precursors of the present invention are as multifunctional cross-linkers for other sol-gel or polymeric systems, and very high surface area materials, i.e., aerogels, when dried via supercritical fluid media. Although the openness of a structure such as star gel precursor 1,  $\text{Si}[\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3]_4$ , suggests that the resulting glass would have open porosity and very high surface area, pore collapse can occur during drying via simple solvent evaporation when the network is sufficiently flexible leading to non-porous coatings or materials.

20        The present invention further comprises a process for the preparation of an inorganic/organic composition of formula (II). To form the composition of formula (II) of the present invention as defined above the alkoxy-silane, acyloxy-silane or halosilane groups of the star gel precursors of formula (I) or formula (III) as previously defined, are hydrolyzed with either water in the presence of a solvent and a catalyst, or one or more strong carboxylic acids, preferably formic acid, optionally in the presence of a solvent and condensed to form a continuous network of silicon-oxygen bonds. The silicon atoms bearing the hydrolyzed groups will be constituents of an infinite network structure via bonds to other silicon atoms through oxygen. Preferred star

gel precursors of formula I or formula III for use in this process include those listed in Table I.

The present invention also further comprises a method for preparing compositions of glasses of  
 5 formula II by combining two or more star gel precursors of the present invention of formula (I) or formula III with each other. If more than one compound of formula (I) or formula III, as defined above, is mixed, the star gel precursors may be represented as  
 10  $X'(SiQ_3)_n' + X''(SiQ_3)_n'' + \dots$ , wherein  $X'$  and  $X''$  are different definitions of  $X$  and  $n'$  and  $n''$  correspond to the definitions of  $X'$  and  $X''$  respectively. The resulting inorganic/organic composition of formula (II), as defined above, will be  $X(SiO_{1.5})_n$ , where  $X = \% X' +$   
 15  $\% X'' + \dots$  and  $n = \text{average of } (\% n' + \% n'' + \dots)$ . Star-derived glasses may also be produced in the presence of a dye such as Rhodamine G to yield an optically useful material.

The process of this invention has the desirable  
 20 feature that no water need be added to the reactants initially and that the steady state water concentration during reaction can be quite small. One of the benefits of this feature is that clear gels can be made readily without a need to use a water-miscible solvent to obtain  
 25 a homogeneous medium. While water is necessary for hydrolysis, a sufficient amount is formed by reaction of strong carboxylic acid with alcohol produced by hydrolysis and by the metathesis reaction:



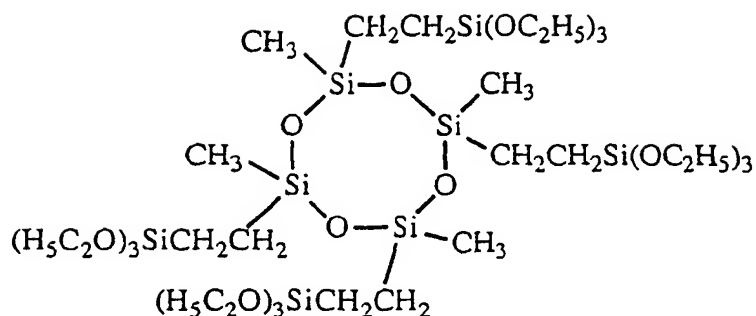
Also, any water which may be present as diluent in strong carboxylic acid can contribute to hydrolysis. Strong carboxylic acid containing at most 20 mol% water  
 35 is preferred for the process of this invention. The

carboxylic acids should have a  $pK_a$  value not higher than about 4.0 and contain 0 to 20 mole % water. Examples of strong carboxylic acids effective in this invention include formic acid, monochloroacetic acid, dichloro-  
5 acetic acid, trifluoroacetic acid and hydroxyacetic acid. Formic acid is the preferred carboxylic acid. After the star gel precursor is mixed with water in the presence of a solvent and a catalyst, or the strong carboxylic acid optionally in the presence of a solvent,  
10 the mixture is maintained at a temperature within the range of about 0-100°C at ambient pressure. The star gel is finally isolated via removal of liquid byproducts and unreacted starting materials to yield a glass.

When the star gel precursors of formula (I) or  
15 formula (III) are hydrolyzed with water in the presence of a solvent and a catalyst, suitable solvents comprise co-solvents for water and the star gel precursor of formula (I) or formula (III) or are miscible with water with an affinity for the star gel precursor of formula  
20 (I) or formula (III), e.g., alcohols, tetrahydrofuran, and acetonitrile. Suitable catalysts comprise Bronsted acids or weak bases where  $pH < 9$ , e.g., hydrogen fluoride, sodium fluoride, sulfuric acid, acetic acid and ammonium hydroxide.

25 Those glasses of formula (II) prepared by drying the gels of formula (II) wherein the X component corresponds to a linear or cyclosiloxane show the greatest flexibility. This flexibility provides a more compliant network structure. The compliance  
30 incorporated into the network allows faster drying rates and imparts toughness to the resultant glasses prepared from gels of formula (II).

Star gel precursor 1,  $Si[CH_2CH_2Si(OC_2H_5)_3]_4$ , star gel precursor 2,



and star gel precursor 3,  $\text{Si}[\text{OSi}(\text{CH}_3)_2\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3]_4$ ,  
 5 in Table 1 are readily soluble in tetrahydrofuran and  
 mixtures of that solvent with water or formic acid. The  
 latter, formic acid, has been developed as a highly  
 effective hydrolytic and condensation agent for  
 tetraalkoxysilanes. Several star gel precursors, e.g.,  
 10 3, can be added directly to formic acid to give a  
 dispersion which rapidly clarifies as the molecule  
 begins to react and silanol groups are generated.  
 Gelation rates with formic acid can be extremely fast in  
 the absence of solvent. Star gel precursor 3,  
 15  $\text{Si}[\text{OSi}(\text{CH}_3)_2\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3]_4$  forms a gel with a 20-fold  
 molar excess of  $\text{HCOOH}$  in 6 minutes. The use of a  
 H-bonding acceptor solvent such as tetrahydrofuran can  
 attenuate the gelation rate up to several orders of  
 magnitude. A less reactive solvent such as methylene  
 20 chloride will give a more rapid gelation rate than a  
 hydrogen bonding solvent. Pure star gel precursors have  
 long shelf life. Hydrolytic reagent and optional  
 solvent and catalyst if water is used are added to  
 initiate the reaction. Gels are clear except for the  
 25 coloration imparted by any residual catalyst. They may  
 be dried into monolithic glassy solids at rates at least  
 five times those which lead to fracture of conventional  
 gels of the same dimensions. Thick films of these  
 glasses can be easily dried without any observed

cracking upon drying. In this manner films that are five times thicker than those derived from sol-gel silica can be made crack free.

Dried samples of glass of formula (II) derived from star gel precursors 1, 2 and 3 as defined above and in Table I, do not show evidence of open porosity when submerged under water. Adsorption isotherm measurements using nitrogen at 77°K also indicated no detectable surface-connected porosity for a sample of the glass derived from star gel precursor 1,  $\text{Si}[\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3]_4$ .

Impact resistance was examined by dropping a 150 g pestle from various heights onto pieces of star-derived and conventional sol-gel glasses of comparable size. The star glasses were able to sustain impacts which invariably fractured their conventional counterparts.

The new classes of compositions of formula (II), in addition to those cited above, are useful as abrasion resistant materials, impact resistant glasses, microporous glasses, interlayer dielectrics for electronic devices, adhesives for glass and other materials, and barrier coatings. Star gel precursors of formula (I), are useful as crosslinking agents for some functionalized organic polymers, coupling agents, or modifiers for alkoxysilane derived sol-gel glasses, other metal alkoxide derived sol-gel glasses, and other star-gel glasses.

The present invention further comprises a method for modifying conventional sol-gel glasses as defined in Brinker, C. J., et al., Sol Gel Science, Academic Press, San Diego, CA (1990), to increase drying rates and lower brittleness comprising combining a precursor of the present invention of formula (I) or of formula (III) with a conventional sol-gel system based on tetraalkoxysilanes or other metal alkoxides; mixing in water with a solvent and a catalyst, or a carboxylic acid, preferably



formic acid, optionally in the presence of a solvent; and drying. The known tetraalkoxysilane or other metal alkoxide is combined with a star gel precursor of the present invention, preferably the star gel precursors include those found in Table I, to generate a homogenous solution. These components are miscible and useable in any proportion, e.g., from 0.1:99.9 to 99.9:to 0.1. Water with a solvent and a catalyst, or a carboxylic acid, preferably formic acid, optionally in the presence of a solvent, is then added with stirring at ambient temperature and pressure to induce gelation. The resulting gel is then dried. Typically drying is at atmospheric pressure and at a temperature of from about 20°C to 150°C. Vacuum up to  $10^{-5}$  torr may be employed. The gelation rate of  $\text{Si}(\text{OC}_2\text{H}_5)_4$  by formic acid can be profoundly influenced by addition of small amounts of a star gel precursor of formula (I) or formula (III). A mixture of  $\text{Si}(\text{OC}_2\text{H}_5)_4$  and  $\text{HCOOH}$  at a molar ratio of 1:3 normally requires 18 hours to gel. Substitution of 10 mole % star gel precursor 1,  $\text{Si}[\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3]_4$ , for  $\text{Si}(\text{OC}_2\text{H}_5)_4$  led to a gelation time of 8 minutes under comparable conditions. Other sol-gel glasses from inorganic alkoxides, for example alkoxides of Al, Zr, V, B, Ti, Nb, Ge, Sn, Ga, In, Cu and Pb can be modified in a similar fashion.

A method of coating a substrate is also provided by the present invention comprising reacting a star gel precursor of formula (I) or formula (III), preferably including those in Table I, with water in the presence of a solvent and a catalyst, or a carboxylic acid, such as formic acid, optionally in the presence of a solvent, such as tetrahydrofuran, dipping the substrate in the resulting mixture, removing the coated substrate from the mixture and drying the coating. Thus the substrate is dipped into the mixture containing the star gel

precursor prior to gelation, and after gelation and drying the substrate is coated with an inorganic/organic composition of formula (II). The star gel precursor of formula (I) or formula (III) or inorganic/organic composition of formula (II) may also be used as an adhesive by coating a substrate and placing another substrate on top of it and applying pressure, optionally accompanied by or followed by heat. Suitable substrates comprise glass, metal and plastic.

- 10 In the examples which follow all star numbers refer to the star gel precursors listed in Table I. All reactions were carried out in a Vacuum Atmospheres Co. dry box under nitrogen. Commercial reagents were distilled prior to use. Triethoxysilane, tetravinylsilane, vinyltriethoxysilane, 1,3,5,7-tetramethylcyclotetrasilane, 1,3,5,7-tetravinyltetramethylcyclotetrasilane, 1,1,3,3-tetravinylldimethyldisiloxane, tetraallyloxysilane, tetrakis(dimethylsiloxy)silane, p-bis-(dimethylsilyl)benzene, bis[p-dimethylsilyl]phenyl-ether, 1,1,3,3-tetramethyldisiloxane, 1,1,4,4-tetramethyldisilethylene, pentamethylcyclopentasiloxane, methyltris(dimethylsiloxy)silane, chlorodimethylvinylsilane, tetraethoxysilane and trichlorosilane were purchased from Huls America Inc., Piscataway, NJ.
- 25 Allyltriethoxysilane; trimethoxysilane; triethoxysilane, 1,3,5-tribromobenzene; 1,2,4-trivinylcyclohexane and tert-butyl lithium (1.7M in pentane) were purchased from Aldrich Chemical Co., Milwaukee, WI. Tetraethoxysilane was purchased from Eastman Kodak, Rochester, NY.
- 30 Platinum divinylsiloxane complex (3-3.5% Pt concentration in xylene, Huls PC072) was obtained from Huls America Inc. and diluted 5:1 by volume (toluene, Pt complex) prior to use. Cobalt carbonyl and  $P(OCH_3)_3$  were obtained from E. I. du Pont de Nemours and Company.
- 35 Toluene was reagent grade and purified by distillation

from lithium aluminum hydride prior to use. Tetrallyl-silane was synthesized by a modification of a published procedure (J. Organomet. Chem., 84(1975), pp. 199-299). 1,3,5-(CH<sub>2</sub>=CH(CH<sub>3</sub>)<sub>2</sub>Si)<sub>3</sub>C<sub>6</sub>H<sub>3</sub> was synthesized by a  
5 modification of a published procedure using CH<sub>2</sub>=CHCH<sub>3</sub>)<sub>2</sub>SiCl instead of Si(OC<sub>2</sub>H<sub>5</sub>)<sub>4</sub> (Macromolecules, 24 (1991), pp. 6863-6866). The preparation of silicon alkoxides (Si-OR) from chlorosilanes (Si-Cl) and alcohol was accomplished according to known procedures (Organo-  
10 silicon Compounds, C. Earborn, Academic Press Inc., NY, 1960, pp. 288-311). Vinylpolyfluoroalkanes CH<sub>2</sub>=CH(CF<sub>2</sub>)<sub>p</sub>CH=CH<sub>2</sub> (p=6,10), CH<sub>2</sub>=CH(CH<sub>2</sub>)<sub>4</sub>(CF<sub>2</sub>)<sub>10</sub>(CH<sub>2</sub>)<sub>4</sub>CH=CH<sub>2</sub>, and ICH<sub>2</sub>CH<sub>2</sub>(CF<sub>2</sub>)<sub>6</sub>CH<sub>2</sub>CH<sub>2</sub>I were obtained pure from E. I. du Pont de Nemours and  
15 Company, Wilmington, DE. Normal purification of the star gel precursors involved flash chromatography on silica gel using hexane as the elutant unless otherwise noted. The silica gel column was treated with Si(OCH<sub>3</sub>)<sub>4</sub> before addition of the star gel precursors. The K<sup>+</sup>IDS  
20 mass spectroscopy experiments were performed on a Finnigan 4615B GC/MS quadrupole mass spectrometer (San Jose, CA). An electron impact source configuration operating at 200°C and a source pressure of 1.0 x 10<sup>-6</sup> Torr was used. The mass spectrometer was  
25 scanned at a rate of about 1000 Daltons/second. All K<sup>+</sup>IDS mass spectral peaks are recorded as sum of the ion plus potassium (M + 39). Proton and carbon NMR were determined in deuterobenzene solvent on a GE model QE-300 instrument. Elemental analyses were performed by  
30 Oneida Research Services Inc., One Halsey Road, Whitesboro, NY.

EXAMPLESEXAMPLE 1

Synthesis and Characterization of Star 1,  
 $\text{Si}[\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3]_4$

5 To a mixture of 55.596 g (0.338 mol) of triethoxy-  
silane and 10 drops (approximately 0.3 ml) of Pt  
catalyst was added 5.219 g (0.038 mol) of tetravinyl-  
silane dropwise over a period of 1 hour. The  
temperature of the reaction mixture was controlled so as  
10 to not exceed 35°C. After the addition, the solution  
was heated to 90°C for 6 hours, then cooled and stirred  
at room temperature for 18 hours. The excess triethoxy-  
silane was removed in vacuo at 60°C. Proton NMR of the  
product showed some residual vinyl groups. An  
15 additional 11.842 g (0.0720 mol) of triethoxysilane and  
4 drops of Pt catalyst was added to the crude mixture  
and heated to 90°C for 6 hours. Cooling to room  
temperature and workup as described above yielded  
26.75 g (88%) of a clear liquid determined to be mostly  
20  $\text{Si}[\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3]_4$ . Purity was found to be 91% by  
K<sup>+</sup>IDS mass spectroscopy and > 75% by supercritical fluid  
chromatography (SFC). K<sup>+</sup>IDS MS (m/e) 831 (M + 39,  
100%), 667 ( $\text{H}_2\text{C}=\text{CH}$ ) $\text{Si}[\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3]_3$  + 39, 8.6%).  
13C NMR( $\text{C}_6\text{D}_6$ ) 3.57 (SiCH<sub>2</sub>), 4.05 (SiCH<sub>2</sub>), 19.0 (CH<sub>3</sub>),  
25 59.0 (SiOCH<sub>2</sub>). Small amounts of -SiCH(CH<sub>3</sub>)Si(OC<sub>2</sub>H<sub>5</sub>)<sub>3</sub>  
groups due to Markovnikov (or reverse-hydrosilation)  
addition (1.0, 9.0 ppm) were observed. Anal: Calcd for  
 $\text{C}_{32}\text{H}_{76}\text{Si}_5\text{O}_{12}$  C, 48.45; H, 9.65; Si, 17.70. Found: C,  
47.72; H, 9.59; Si, 17.37.

30

EXAMPLE 2

Synthesis and Characterization of Star 2,  
 $((\text{CH}_3)_4(\text{C}_2\text{H}_5\text{O})_3\text{SiCH}_2\text{CH}_2)_4(\text{SiO})_4$

To a stirred mixture of 2.85 g (0.0083 mol) of  
 $((\text{CH}_3)(\text{CH}_2=\text{CH})\text{SiO})_4$  and 8.15 g (0.0496 mol) of triethoxy-  
35 silane was added 14 drops (ca 0.4 ml) of Pt catalyst.

The resulting solution was heated to 100°C for 2.5 hours, cooled and stirred at room temperature for 18 hours. The excess  $\text{HSi}(\text{OC}_2\text{H}_5)_3$  was removed in vacuo and workup as described previously yielded a clear liquid identified as  $((\text{CH}_3)((\text{C}_2\text{H}_5\text{O})_3\text{SiCH}_2\text{CH}_2)\text{SiO})_4$ . Impurities were mainly the di- and tri-substituted products as noted by  $\text{K}^+\text{IDS MS}$ .  $^{13}\text{C NMR}(\text{C}_6\text{D}_6)$  -0.901  $((\text{CH}_3)\text{Si})$ , 2.98  $(\text{SiCH}_2)$ , 9.30  $(\text{SiCH}_2)$ , 19.01  $(\text{CH}_3)$ , 58.99  $(\text{SiOCH}_2)$ .  $\text{K}^+\text{IDS MS (m/e)}$  1039  $(\text{M} + 39, 100 \%)$ , 875  $(3\text{-arm product} + 39, 52 \%)$ , 711  $(2\text{-arm product} + 39, 3 \%)$ . Anal. Calcd for  $\text{C}_{36}\text{H}_{88}\text{Si}_8\text{O}_{16}$ : C, 43.16; H, 8.85. Found: C, 42.12; H, 8.65.

#### EXAMPLE 3

##### Synthesis and Characterization of Star 3,

15  $\text{Si}(\text{OSi}(\text{CH}_3)_2\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3)_4$

To a stirred mixture of 3.039 g (0.0092 mol) of  $\text{Si}(\text{OSi}(\text{CH}_3)_2\text{H})_4$  and 10.024 g (0.05278 mol) of vinyl-triethoxysilane was added 14 drops (ca 0.4 ml) of Pt catalyst. The resulting solution was stirred for 20 2 hours, heated to 90°C for 4 hours, cooled and stirred at room temperature for 18 hours. The excess  $(\text{CH}_2=\text{CH})\text{Si}(\text{OC}_2\text{H}_5)_3$  was removed in vacuo and workup as described previously yielded a clear liquid identified as  $\text{Si}(\text{OSi}(\text{CH}_3)_2\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3)_4$ . Impurity levels were 25 on the order of 5% or less.  $^{13}\text{C NMR}(\text{C}_6\text{D}_6)$  -0.43  $((\text{CH}_3)\text{Si})$ , 3.19  $(\text{SiCH}_2)$ , 10.17  $(\text{SiCH}_2)$ , 19.09  $(\text{CH}_3)$ , 58.86  $(\text{SiOCH}_2)$ .  $\text{K}^+\text{IDS MS (m/e)}$  1128  $(\text{M} + 39, 100 \%)$ . Anal. Calcd for  $\text{C}_{40}\text{H}_{100}\text{Si}_9\text{O}_{16}$ : C, 44.08; H, 9.25; Si, 23.19. Found: C, 44.66; H, 9.31; Si, 22.46.

30 EXAMPLE 4

##### Synthesis and Characterization of Star 4,

$\text{Si}(\text{OCH}_2\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3)_4$

A mixture of 3.12 g (0.0118 mol) of  $\text{Si}(\text{OCH}_2\text{CH}=\text{CH}_2)_4$ , 11.028 g (0.0671 mol) of  $\text{HSi}(\text{OC}_2\text{H}_5)_3$  and 35 14 drops of Pt catalyst was stirred at 25°C for 2 hours

and then heated to 90°C for 4 hours. The solution was cooled, and the volatiles removed in vacuo. Workup as described above yielded 6.49 g of a clear liquid identified as  $\text{Si}(\text{OCH}_2\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3)_4$ .  $^{13}\text{C}$  NMR ( $\text{C}_6\text{D}_6$ ) 7.39 ( $\text{SiCH}_2$ ), 18.97 ( $\text{CH}_3$ ), 26.93 ( $\text{CH}_2$ ), 58.89 ( $\text{SiOCH}_2\text{CH}_3$ ), 66.55 ( $\text{SiCH}_2$ ).  $\text{K}^+\text{IDS}$  MS ( $m/e$ ) 952 ( $M + 39$ , 100%). Anal. Calcd for  $\text{C}_{36}\text{H}_{84}\text{Si}_5\text{O}_{16}$ : C, 47.37; H, 9.27; Si, 15.37. Found: C, 46.32; H, 9.10; Si, 16.06.

#### EXAMPLE 5

Synthesis and Characterization of Star 5,  
 $((\text{C}_2\text{H}_5\text{O})_3\text{SiCH}_2\text{CH}_2)_2(\text{CH}_3)\text{SiOSi}(\text{CH}_3)(\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3)_2$   
 To a stirred solution of 2.12 g (0.0101 mol) of  $((\text{CH}_2=\text{CH})_2(\text{CH}_3)\text{Si})_2\text{O}$  and 14 drops of Pt catalyst was added 7.93 g (0.0483 mol) of  $\text{HSi}(\text{OC}_2\text{H}_5)_3$  via a syringe over a period of 1 hour. The mixture was heated to 90°C for 4 hours, then cooled to room temperature. The volatiles were removed in vacuo and the reaction product was worked up as described above yielding 3.80 g of a liquid identified as  
 $((\text{C}_2\text{H}_5\text{O})_3\text{SiCH}_2\text{CH}_2)_2(\text{CH}_3)\text{SiOSi}(\text{CH}_3)(\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3)_2$ .  
 Small amounts of impurities (< 5%) were noted in the NMR corresponding to the tri-substituted product.  $^{13}\text{C}$  NMR ( $\text{C}_6\text{D}_6$ ) -2.01 ( $\text{CH}_3\text{Si}$ ), 3.27 ( $\text{SiCH}_2$ ), 8.32 ( $\text{SiCH}_2$ ), 19.02 ( $\text{CH}_3$ ), 58.92 ( $\text{SiOCH}_2$ ).  $\text{K}^+\text{IDS}$  MS ( $m/e$ ) 905 ( $M + 39$ , 100%), 743 (3-arm product + 39, 25 %). Anal. Calcd for  $\text{C}_{34}\text{H}_{82}\text{Si}_6\text{O}_{13}$ : C, 47.07; H, 9.53; Si, 19.43. Found: C, 45.69; H, 9.40; Si, 19.40.

#### EXAMPLE 6

Synthesis and Characterization of Star 6,  
 $(\text{C}_2\text{H}_5\text{O})_3\text{SiCH}_2\text{CH}_2(\text{CH}_3)_2\text{SiC}_6\text{H}_4\text{Si}(\text{CH}_3)_2\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3$   
 A mixture of 2.91 g (0.0150 mol) of  $\text{H}(\text{CH}_3)_2\text{C}_6\text{H}_4(\text{CH}_3)_2\text{H}$ , 10.03 g (0.0527 mol) of  $(\text{CH}_2=\text{CH})\text{Si}(\text{OC}_2\text{H}_5)_3$  and 14 drops of Pt catalyst was stirred at 25°C for 2 hours and then heated to 90°C for 4 hours. The solution was cooled, and the volatiles

removed in vacuo. Workup as described above yielded 11.6 g of a liquid identified as Star 6.  $^{13}\text{C}$  NMR( $\text{C}_6\text{D}_6$ ) -3.11 ( $\text{SiCH}_3$ ), 3.85 ( $\text{SiCH}_2$ ), 7.80 ( $\text{SiCH}_2$ ), 19.0 ( $\text{CH}_3$ ), 26.93 ( $\text{CH}_2$ ), 58.9 ( $\text{SiOCH}_2\text{CH}_3$ ), 133.7, 136.6, 140.2

- 5 (aromatics).  $\text{K}^+\text{IDS}$  MS ( $m/e$ ) 614 ( $M + 39$ , 100 %). Anal. Calcd for  $\text{C}_{26}\text{H}_{54}\text{Si}_4\text{O}_6$ : C, 54.31; H, 9.46. Found: C, 53.53; H, 9.40.

#### EXAMPLE 7

Synthesis and Characterization of Star 7,

- 10  $(\text{C}_2\text{H}_5\text{O})_3\text{SiCH}_2\text{CH}_2(\text{CH}_3)_2\text{SiC}_6\text{H}_4\text{OC}_6\text{H}_4\text{Si}(\text{CH}_3)_2\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3$   
A mixture of 4.32 g (0.0151 mol) of  $(\text{H}(\text{CH}_3)_2\text{C}_6\text{H}_4)_2\text{O}$ , 10.02 g (0.0526 mol) of  $(\text{CH}_2=\text{CH})\text{Si}(\text{OC}_2\text{H}_5)_3$  and 14 drops of Pt catalyst was stirred at  $25^\circ\text{C}$  for 2 hours and then heated to  $90^\circ\text{C}$  for 4 hours. The solution was cooled, 15 and the volatiles removed in vacuo. The resulting brown liquid was stirred over activated charcoal and filtered yielding 3.59 g of a clear liquid identified as Star 7.  $^{13}\text{C}$  NMR( $\text{C}_6\text{D}_6$ ) -2.93 ( $\text{SiCH}_3$ ), 3.89 ( $\text{SiCH}_2$ ), 7.99 ( $\text{SiCH}_2$ ), 19.06 ( $\text{CH}_3$ ), 58.94 ( $\text{SiOCH}_2\text{CH}_3$ ), 119.2, 133.8, 136.0, 20 158.9 (aromatics). Anal. Calcd for  $\text{C}_{32}\text{H}_{58}\text{Si}_4\text{O}_7$ : C, 57.61; H, 8.76. Found: C, 57.03; H, 8.77.

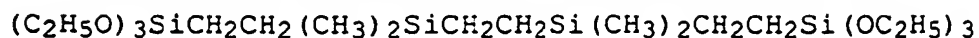
#### EXAMPLE 8

Synthesis and Characterization of Star 8,

- 25  $(\text{C}_2\text{H}_5\text{O})_3\text{SiCH}_2\text{CH}_2(\text{CH}_3)_2\text{SiOSi}(\text{CH}_3)_2\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3$   
A solution consisting of 2.01 g (0.0150 mol) of  $(\text{H}(\text{CH}_3)_2\text{Si})_2\text{O}$ , 10.01 g (0.0526 mol) of  $(\text{CH}_2=\text{CH})\text{Si}(\text{OC}_2\text{H}_5)_3$  and 14 drops of Pt catalyst was stirred at  $25^\circ\text{C}$  for 2 hours and then heated to  $90^\circ\text{C}$  for 4 hours. The solution was cooled, and the volatiles removed in vacuo. 30 Workup as described above yielded 10.67 g of a clear liquid identified as Star 8. Small impurities (about 5%) were noted in the NMR spectra.  $^{13}\text{C}$  NMR( $\text{C}_6\text{D}_6$ ) 0.22 ( $\text{SiCH}_3$ ), 3.24 ( $\text{SiCH}_2$ ), 10.41 ( $\text{SiCH}_2$ ), 18.88 ( $\text{CH}_3$ ), 58.95 ( $\text{SiOCH}_2\text{CH}_3$ ).  $\text{K}^+\text{IDS}$  MS ( $m/e$ ) 553 ( $M + 39$ , 45 %)

EXAMPLE 9

## Synthesis and Characterization of Star 9,



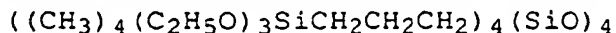
To a mixture consisting of 10.50 g (0.0552 mol) of  
5  $(\text{CH}_2=\text{CH})\text{Si}(\text{OC}_2\text{H}_5)_3$  and 14 drops of Pt catalyst was added  
3.72 g (0.0254 mol) of  $(\text{H}(\text{CH}_3)_2\text{SiCH}_2)_2$  over a 30 min  
period. The temperature was kept around 35°C during the  
addition. The mixture was then heated to 90°C for  
6 hours. After cooling, the volatiles removed in vacuo  
10 yielding 10.39 g of a clear liquid identified as Star 9.  
Some minor impurities were noted in the NMR spectra.  
 $^{13}\text{C}$  NMR( $\text{C}_6\text{D}_6$ ) -3.89 (SiCH<sub>3</sub>), 3.82 (SiCH<sub>2</sub>), 6.71 (SiCH<sub>2</sub>),  
7.46 (SiCH<sub>2</sub>), 19.07 (CH<sub>3</sub>), 58.95 (SiO CH<sub>2</sub> CH<sub>3</sub>). K<sup>+</sup>IDS MS  
(m/e) 565 (M + 39, 100%). Anal. Calcd for  $\text{C}_{22}\text{H}_{54}\text{Si}_4\text{O}_6$ :  
15 C, 50.14; H, 10.33. Found: C, 50.10; H, 10.35.

The reaction was performed in a manner similar to  
Star 9 using 10.03 g (0.0527 mol) of  $(\text{CH}_2=\text{CH})\text{Si}(\text{OC}_2\text{H}_5)_3$ ,  
2.89 g (0.0096 mol) of  $((\text{CH}_3)(\text{H})\text{SiO})_5$  and 14 drops of Pt  
catalyst. Workup yielded 8.66 g of a clear liquid  
20 identified as Star 10.  $^{13}\text{C}$  NMR( $\text{C}_6\text{D}_6$ ) -0.75 (SiCH<sub>3</sub>), 3.12  
(SiCH<sub>2</sub>), 9.48 (SiCH<sub>2</sub>), 19.05 (CH<sub>3</sub>), 58.95 (SiO CH<sub>2</sub> CH<sub>3</sub>).  
K<sup>+</sup>IDS MS (m/e) 1290 (M + 39, 100 %). Anal. Calcd for  
 $\text{C}_{45}\text{H}_{110}\text{Si}_{10}\text{O}_{20}$ : C, 43.16; H, 8.85. Found: C, 43.15; H,  
8.79.

25

EXAMPLE 11

## Synthesis and Characterization of Star 12,



The reaction was performed in a manner similar to  
Star 9 using 10.04 g (0.0491 mol) of  
30  $(\text{CH}_2=\text{CHCH}_2)\text{Si}(\text{OC}_2\text{H}_5)_3$ , 2.32 g (0.0097 mol) of  
 $((\text{CH}_3)(\text{H})\text{SiO})_4$  and 14 drops of Pt catalyst. Workup  
yielded 7.74 g of a liquid identified as Star 12.  $^{13}\text{C}$   
NMR( $\text{C}_6\text{D}_6$ ) 0.131 (SiCH<sub>3</sub>), 15.57 (SiCH<sub>2</sub>), 17.64 (SiCH<sub>2</sub>),  
19.03 (CH<sub>3</sub>), 22.07 (CH<sub>2</sub>), 58.79 (SiO CH<sub>2</sub> CH<sub>3</sub>). K<sup>+</sup>IDS MS



(m/e) 1095 (M + 39, 100%). Anal. Calcd for  $C_{40}H_{96}Si_8O_{16}$ : C, 45.42; H, 9.15. Found: C, 46.35; H, 9.26.

#### EXAMPLE 12

##### Synthesis and Characterization of Star 13,

5  $((CH_3)_5(C_2H_5O)_3SiCH_2CH_2CH_2)_5(SiO)_5$

The reaction was performed in a manner similar to Star 9 using 8.49 g (0.0416 mol) of  $(CH_2=CHCH_2)Si(OC_2H_5)_3$ , 2.45 g (0.0082 mol) of  $((CH_3)(H)SiO)_5$  and 14 drops of Pt catalyst. Workup  
10 yielded 5.94 g of a liquid identified as Star 12.  $^{13}C$  NMR( $C_6D_6$ ) 0.269 (SiCH<sub>3</sub>), 15.74 (SiCH<sub>2</sub>), 17.79 (SiCH<sub>2</sub>), 19.04 (CH<sub>3</sub>), 22.40 (CH<sub>2</sub>), 58.83 (SiOCH<sub>2</sub>CH<sub>3</sub>).  $K^+IDS$  MS  
(m/e) 1359 (M + 39, 100%). Anal. Calcd for  $C_{50}H_{120}Si_{10}O_{20}$ : C, 45.42; H, 9.15. Found: C, 46.41; H,  
15 9.23.

#### EXAMPLE 13

##### Synthesis and Characterization of Star 14,

$Si(OSi(CH_3)_2CH_2CH_2CH_2Si(OC_2H_5)_3)_4$

To a stirred mixture of 10.04 g (0.0491 mol) of  
20 allyltriethoxysilane and 14 drops (ca 0.4 ml) of Pt catalyst was added 3.17 g (0.0096 mol) of  $Si(OSi(CH_3)_2H)_4$  over a period of 1 hour. The resulting solution was heated to 90°C for 4 hours, cooled and stirred at room temperature for 18 hours. The excess  
25  $(CH_2=CHCH_2)Si(OC_2H_5)_3$  was removed in vacuo and workup as described previously yielded 8.84 g of a liquid identified as  $Si(OSi(CH_3)_2CH_2CH_2CH_2Si(OC_2H_5)_3)_4$ .  $^{13}C$  NMR( $C_6D_6$ ) -0.14 ((CH<sub>3</sub>)Si), 15.04 (SiCH<sub>2</sub>), 16.74 (SiCH<sub>2</sub>), 17.80 (CH<sub>3</sub>), 22.13 (-CH<sub>2</sub>-), 58.86 (SiOCH<sub>2</sub>).  $K^+IDS$  MS  
30 (m/e) 1183 (M + 39, 100%). Anal. Calcd for  $C_{44}H_{108}Si_5O_{16}$ : C, 46.11; H, 9.50. Found: C, 46.28; H, 9.55.

EXAMPLE 14

Synthesis and Characterization of Star 15,  
 $(C_2H_5O)_3SiCH_2CH_2CH_2(CH_3)_2SiC_6H_4Si(CH_3)_2CH_2CH_2CH_2Si(OC_2H_5)_3$

To a stirred mixture of 9.56 g (0.0468 mol) of  
5 allyltriethoxysilane and 14 drops (ca 0.4 ml) of Pt  
catalyst was added 2.94 g (0.0151 mol) of  
 $H(CH_3)_2SiC_6H_4Si(CH_3)_2H$  over a period of 30 min. The  
resulting solution was heated to 90°C for 4 hours,  
cooled and stirred at room temperature for 18 hours.  
10 The excess  $(CH_2=CHCH_2)Si(OC_2H_5)_3$  was removed in vacuo and  
workup as described previously yielded 9.63 g of a  
liquid identified as Star 15.  $^{13}C$  NMR( $C_6D_6$ ) -2.43  
( $(CH_3)Si$ ), 15.96 ( $SiCH_2$ ), 18.53 ( $SiCH_2$ ), 19.00 ( $CH_3$ ),  
20.42 ( $-CH_2-$ ), 58.82 ( $SiOCH_2$ ).  $K^+IDS$  MS (m/e) 641 (M +  
15 39, 100%).

EXAMPLE 15

Synthesis and Characterization of Star 17,  
 $(CH_3)_3Si(OSi(CH_3)_2CH_2CH_2Si(OC_2H_5)_3)_3$

A stirred mixture of 10.01 g (0.0609 mol) of  
20 triethoxysilane, 14 drops (ca 0.4 ml) of Pt catalyst and  
3.35 g (0.0123 mol) of  $(CH_3)_3Si(OSi(CH_3)_2CH=CH_2)_3$  was  
stirred at 25°C for 2 hours and then heated to 90°C for  
4 hours, cooled and stirred at room temperature for  
18 hours. The excess  $HSi(OC_2H_5)_3$  was removed in vacuo  
25 yielding 15.46 g of a clear liquid identified as  
 $(CH_3)_3Si(OSi(CH_3)_2CH_2CH_2Si(OC_2H_5)_3)_3$ .  $^{13}C$  NMR( $C_6D_6$ ) 0.11  
( $(CH_3)Si$ ), 3.16 ( $SiCH_2$ ), 10.30 ( $SiCH_2$ ), 18.82 ( $CH_3$ ),  
58.76 ( $SiOCH_2$ ). Anal. Calcd for  $C_{31}H_{78}Si_7O_{12}$ : C, 44.35;  
H, 9.36. Found: C, 44.99; H, 9.46

30

EXAMPLE 16

Synthesis and Characterization of Star 18,  
 $Si(CH_2CH_2Si(OCH_3)_3)_4$

To a stirred mixture of 2.21 g (0.0162 mol) of  
tetravinylsilane and 7 drops (ca 0.4 ml) of Pt catalyst  
35 was added 12.016 g (0.0096 mol) of trimethoxysilane over

a period of 1 hour. The resulting solution was heated to 90°C for 4 hours, cooled and stirred at room temperature for 18 hours. The excess  $\text{HSi}[\text{O}(\text{CH}_3)]_3$  was removed in vacuo yielding 9.78 g of a liquid identified as  $\text{Si}(\text{CH}_2\text{CH}_2\text{Si}(\text{OCH}_3)_3)_4$ . Column chromatography of the Star product resulted in partial hydrolysis of Si-OCH<sub>3</sub> groups. Some Markovnikov addition products, similar to those observed in Star 1, were also seen in the NMR. <sup>13</sup>C NMR(C<sub>6</sub>D<sub>6</sub>) 2.69 (SiCH<sub>2</sub>), 3.34 (SiCH<sub>2</sub>), 50.8 (SiOCH<sub>3</sub>). K<sup>+</sup>IDS MS (m/e) 663 (M + 39, 90%), 617 ((CH<sub>3</sub>)<sub>2</sub>O + 39, 100%).

#### EXAMPLE 17

Gel from Star 1 with water/ethanol

1.868 g Star 1 was dissolved in 2.42 g anhydrous ethanol to give a homogeneous solution. 0.510 g of water containing 0.040 g 0.1 N HCl was added to this solution over a period of two minutes with moderate stirring. The ratio of water to Si(OR) groups was 1.00. The solution remained clear and formed a clear gel after 23 hours. The wet gel was dried initially at room temperature and atmospheric pressure, then at a temperature of 120°C and a pressure of 10<sup>-5</sup> torr. The dry gel was pulverized and subjected to pore size analysis by the BET method adsorption analysis using nitrogen at -196°C in accordance with ASTM standard C1069-86. The sample did not show a significant weight gain due to nitrogen adsorption after 4 hours. Small pieces of the dried gel were submerged in water and observed under a microscope. No evidence of fracture or gas liberation was observed. Taken together, these observations indicate the sample did not possess open porosity.

EXAMPLE 18

Gel from Star 3 with HCOOH

1.562 g Star 3 was added to 1.86 g 96% formic acid with stirring. The mixture formed a clear solution after several seconds. The ratio of acid to Si(OR) groups was 2.26. The solution was transferred into a polyethylene vial and allowed to stand at room temperature. The sample formed a transparent gel after 13 min. The contents of the vial were allowed to dry via slow evaporation of the liquid component of the gel. The material was fully dried after two weeks and had formed an intact smooth right cylinder which was translucent. The dried gel remained intact after impact from a 150 g pestle dropped from a heights of 2-3 cm. Comparably sized pieces of conventional sol-gel glasses were consistently fractured from the same impact exposure.

EXAMPLE 19

Attenuation of Gelation Rate for Star 3

with Tetrahydrofuran

1.183 g Star 3 was dissolved in 3.20 g tetrahydrofuran (THF). 0.526 g 96% formic acid was added to above solution with stirring. The ratio of acid to Si(OR) groups was 0.842. The solution was transferred into a fluoropolymer vial and allowed to stand at room temperature. The sample formed a transparent gel after 19 days.

EXAMPLE 20

Dip and Flow coats from Star 1

HCOOH/Tetrahydrofuran

3.36 g Star 1 was combined with 12.37 g tetrahydrofuran and 2.35 g formic acid per Example 4. The ratio of acid to Si(OR) groups was 0.96. The solution (which gelled in four hours) was used to prepare coatings on glass slides via dip and flow coating techniques.

Coatings made three hours after the reactants were mixed were transparent and crack-free. The flow coating was shown by surface profilometry to be 2.5 micrometers in thickness; the dip coating was 0.5 micrometers in thickness.

#### EXAMPLE 21

Gel from Star 1 with tetraethoxysilane in  
HCOOH; Gel rate enhancement

0.865 Star 1 was combined with 2.04 g tetraethoxy-  
silane to give a homogeneous solution which was added to  
1.53 g HCOOH with stirring. The resultant solution  
gelled in 8.3 minutes. The Star was 10 mole % of the  
total silanes present. The ratio of acid to silanes was  
2.93. At the same molar ratio of HCOOH/silane, pure  
tetraethoxysilane requires ca. 18 hours for gelation.

#### EXAMPLE 22

High Surface Area Gel from Star 1  
with tetraethoxysilane

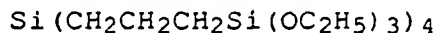
1.064 g Star 1 combined with 2.492 g tetraethoxy-  
silane to give a homogeneous solution which was added to  
3.161 g HCOOH with stirring. The resultant solution  
gelled in 1.5 minutes. The Star was 10 mole % of the  
total silanes present. At the same molar ratio of  
HCOOH/silane (4.96), pure tetraethoxysilane requires ca.  
2 hours for gelation. The wet gel was dried under  
vacuum at 60°C within minutes of its preparation, then  
at a temperature of 120°C and a pressure of  $10^{-5}$  torr.  
The dry gel was then subjected to porosity analysis per  
the procedure in Example 17. The surface area was  
determined to be 629 m<sup>2</sup>/g, with an average pore size of  
approximately 2.0 nanometers. The surface area value is  
higher than those observed by this procedure for gels  
made from tetraethoxysilane without the star gel  
precursor.

EXAMPLE 23

Gel from star 10 with HCOOH/methylene chloride  
1.92 g of star gel precursor 10 was dissolved in  
4.05 g reagent grade methylene chloride to give a  
5 homogeneous solution. 0.785 g of 96% formic acid was  
added dropwise to this solution over a period of one  
minute with moderate stirring. The ratio of acid to  
Si(OR) groups was 0.71. The resultant solution remained  
clear and later formed a clear gel on standing  
10 overnight. Several drops of the solution were placed  
between two 25 x 75 mm glass microscope slides so as to  
form a thin continuous layer between the slides. After  
several hours, the slides were firmly bonded together  
and could not be separated or moved relative to each  
15 other by moderate amounts of force, illustrating the  
adhesive nature of the gel.

EXAMPLE 24

## Synthesis and Characterization of Star 11

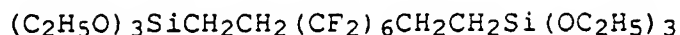


20 { (a); m = 0; k = 0; h = 0; all R's = H }

To a stirred solution of 5.28 g (0.0321 mol) of  
HSi(OC<sub>2</sub>H<sub>5</sub>)<sub>3</sub> and 5 drops of Pt catalyst solution in 20 mL  
of hexane was added 0.626 g (0.0033 mol) of  
Si(CH<sub>2</sub>CH=CH<sub>2</sub>)<sub>4</sub> over a two minute period. The mixture was  
25 refluxed for 1 hr and stirred at 25°C for 60 hr. The  
unreacted volatiles were removed in vacuo, and the crude  
mixture was worked up as described previously yielding  
1.78 g (64%) of Si(CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Si(OC<sub>2</sub>H<sub>5</sub>)<sub>3</sub>)<sub>4</sub>. <sup>13</sup>C NMR(C<sub>6</sub>D<sub>6</sub>)  
16.46, 17.89, 18.65 (SiCH<sub>2</sub>), 19.05 (CH<sub>3</sub>), 58.83  
30 (SiOCH<sub>2</sub>CH<sub>3</sub>). K+IDS MS (m/e) (M+39, 100%).

EXAMPLE 25

## Synthesis and Characterization of Star 16



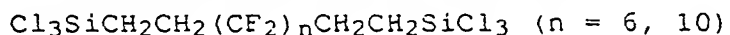
((j); p = 6; k = 0, h = 0, all R's = H)

- 5 To a stirred mixture of 0.041 g (0.120 mmol) of  $\text{Co}_2(\text{CO})_8$ , 0.029 g (0.234 mmol) of  $\text{P}(\text{OCH}_3)_3$  in 1 mL of toluene was added 2.48 g (6.503 mmol) of  $\text{CH}_2=\text{CH}(\text{CF}_2)_6\text{CH}=\text{CH}_2$  and 5.34 g (32.5 mmol) of  $\text{HSi}(\text{OC}_2\text{H}_5)_3$ . The mixture was stirred at room temperature for 4 days
- 10 and an extra 3.22 g of  $\text{HSi}(\text{OC}_2\text{H}_5)_3$  was added to ensure completion of the reaction. After stirring for 11 days the solution was heated at 60°C for 6 hr; and stirred at room temperature for another 7 days until there was no remaining vinyl groups observed in the NMR. Standard
- 15 workup provided 5.19 g of a dark brown solution. The remaining color was removed by the addition of activated charcoal. The ratio of  $\text{CH}_2=\text{CH}(\text{CF}_2)_6\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3$  to  $(\text{C}_2\text{H}_5\text{O})_3\text{SiCH}_2\text{CH}_2(\text{CF}_2)_6\text{CH}_2\text{CH}_2\text{Si}(\text{OC}_2\text{H}_5)_3$  in the product mixture was found to be 74% to 26%. K+IDS MS (m/e) 557
- 20 (monosubstituted, M + 39, 100%), 721 (M+39, 65%). GC/MS (CI-isobutane) exact mass for  $\text{C}_{16}\text{H}_{22}\text{O}_3\text{SiF}_{12} + \text{H}$ , calcd m/e 519.1225, found m/e 519.1263, exact mass for  $\text{C}_{22}\text{H}_{38}\text{O}_6\text{Si}_2\text{F}_{12} + \text{H}$ , calcd m/e 683.2093, found m/e 683.2144.

25

EXAMPLE 26

## Synthesis and Characterization of



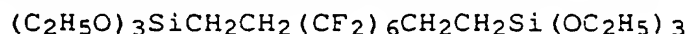
((j); p = 6, 10; k = 0, h = 0, all R's = H; Q = Cl)

- To a 10 mL pressure vessel was added 1.28 g
- 30 (3.36 mmol) of  $\text{CH}_2=\text{CH}(\text{CF}_2)_6\text{CH}=\text{CH}_2$ , 1.32 mL (13.08 mmol) of  $\text{HSiCl}_3$  and one drop of Pt catalyst. The reaction vessel was sealed and heated at 100°C for 48 hr. The vessel was cooled, and the excess  $\text{HSiCl}_3$  was removed in vacuo leaving a 1.622 g (77% yield) of a white solid.
- 35 GC analysis showed that the solid is a single compound.

1H NMR ( $C_6D_6$ ) 1.05 (m, 2H,  $SiCH_2$ ), 1.98 (m, 2H,  $SiCH_2$ ). The preparation of  $Cl_3SiCH_2CH_2(CF_2)_{10}CH_2CH_2SiCl_3$  was performed in a similar manner using 3.66 g (6.66 mmol) of  $CH_2=CH(CF_2)_6CH=CH_2$ , 2.71 g (20.02 mmol) of  $HSiCl_3$ . A temperature of 120°C for 48 hr was needed to ensure complete reaction. Workup yielded 1.63 g (30%) of  $Cl_3SiCH_2CH_2(CF_2)_{10}CH_2CH_2SiCl_3$  as the only product. 1H NMR ( $C_6D_6$ ) 1.02 (m, 2H,  $SiCH_2$ ), 1.91 (m, 2H,  $SiCH_2$ ). The conversion of  $Cl_3SiCH_2CH_2(CF_2)_nCH_2CH_2SiCl_3$  to  $(C_2H_5O)_3SiCH_2CH_2(CF_2)_nCH_2CH_2Si(OC_2H_5)_3$  can be accomplished by known literature methods using ethanol.

EXAMPLE 27

## Synthesis and Characterization of Star 16

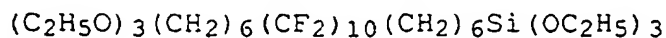


{(j); p = 6, k = 0, h = 0, all R's = H}

A stirred solution containing 0.259 g (0.424 mmol) of  $[ICH_2CH_2(CF_2)_3]_2$  dissolved in 10 mL of ether was cooled to -78°C. To this was added 1.01 mL (1.71 mmol) of tert-butyl lithium. The resulting mixture was stirred for 1.5 h and 1.90 mL (8.518 mmol) of  $Si(OC_2H_5)_4$  was added. The mixture was warmed to room temperature and stirred for 90 hr. The volatiles were removed in vacuo yielding 0.054 g (18%) of a brown residue. GC/MS (CI-isobutane) shows the residue is mainly  $(C_2H_5O)_3SiCH_2CH_2(CF_2)_6CH_2CH_2Si(OC_2H_5)_3$  (calcd m/e/ 683.2156, found  $(C_2H_5O)_3SiCH_2CH_2(CF_2)_6CH_2CH_2Si(OC_2H_5)_3$  (But)).

EXAMPLE 28

## Synthesis and Characterization of Star 21



{(j); p = 10, k = 4, h = 0, all R's = H}

A solution consisting of 5.00 g (7.51 mmol) of  $[CH_2=CH(CH_2)_4(CF_2)_5]_2$ , 5.55 g (33.8 mmol) of  $HSi(OC_2H_5)_3$ , 10 drops of Pt catalyst in 20 mL of toluene was heated at 90°C for 6 hr, and stirred at room temperature for



10 hr. The volatiles were removed in vacuo and normal workup provided 6.45 g (86%) of  $(\text{C}_2\text{H}_5\text{O})_3\text{Si}(\text{CH}_2)_6(\text{CF}_2)_{10}(\text{CH}_2)_6\text{Si}(\text{OC}_2\text{H}_5)_3$  as an off-white waxy solid.  $^{13}\text{C}$  NMR( $\text{C}_6\text{D}_6$ ) 11.50, 20.71, 23.53, 29.27, 33.12 ( $\text{CH}_2$ ), 31.58 (t,  $\text{CH}_2\text{CF}_2$ ,  $2J(\text{C}-\text{F}) = 22$  Hz), 18.99 ( $\text{SiOCH}_2\text{CH}_3$ ), 58.91 ( $\text{SiOCH}_2\text{CH}_3$ ). K+IDS MS (m/e) 1033 (M +39, 100%).

#### EXAMPLE 29

##### Synthesis and Characterization of Star 19

10  $1,3,5-((\text{C}_2\text{H}_5\text{O})_3\text{SiCH}_2\text{CH}_2(\text{CH}_3)_2\text{Si})_3\text{C}_6\text{H}_3$   
{Z =  $\text{CH}_3$ ; k = 0, h = 0, all R's = H}  
To 3.01 g (9.12 mmol) of  $1,3,5-((\text{C}_2\text{H}_5\text{O})_3\text{SiCH}_2\text{CH}_2(\text{CH}_3)_2\text{Si})_3\text{C}_6\text{H}_3$  and eight drops of Pt catalyst was added 5.54 g (33.7 mmol) of  $\text{HSi}(\text{OC}_2\text{H}_5)_3$ .  
15 The resulting mixture was heated to  $90^\circ\text{C}$  for 6 hr and stirred at room temperature for 16 hr. Standard workup provided 5.83 g (78%) of  $1,3,5-((\text{C}_2\text{H}_5\text{O})_3\text{SiCH}_2\text{CH}_2(\text{CH}_3)_2\text{Si})_3\text{C}_6\text{H}_3$  as the sole product.  $^{13}\text{C}$  NMR( $\text{C}_6\text{D}_6$ ) -2.91 ( $\text{CH}_3\text{Si}$ ), 4.05 ( $\text{CH}_2$ ), 7.99 ( $\text{CH}_2$ ),  
20 19.03 ( $\text{SiOCH}_2\text{CH}_3$ ), 58.97 ( $\text{SiOCH}_2\text{CH}_3$ ), 137.96, 140.32 (aromatic). K+IDS MS (m/e) 861 (M +39, 100%).

#### EXAMPLE 30

##### Synthesis and Characterization of Star 20

25  $1,2,4-((\text{C}_2\text{H}_5\text{O})_3\text{SiCH}_2\text{CH}_2)_3\text{C}_6\text{H}_3$   
{k = 0, h = 0, all R's = H}  
A mixture containing 3.206 g (0.0198 mol) of  $1,2,4$ -trivinylcyclohexane, 26.32 g (0.160 mol) of  $\text{HSi}(\text{OC}_2\text{H}_5)_3$  and 10 drops of Pt catalyst was stirred at room temperature. Oxygen was bubbled through the  
30 solution for 5 min., and then the solution was heated to reflux for 7 hr, cooled and stirred at room temperature for 16 hr. Standard workup provided 11.68 g (90%) of  $1,2,4-((\text{C}_2\text{H}_5\text{O})_3\text{SiCH}_2\text{CH}_2)_3\text{C}_6\text{H}_3$  as a clear liquid.  $^{13}\text{C}$  NMR( $\text{C}_6\text{D}_6$ ) 5.98 to 42.72 (many peaks;  $\text{CH}_2$ , CH), 18.18  
35 ( $\text{SiOCH}_2\text{CH}_3$ ), 17.97 (disubst,  $\text{SiOCH}_2\text{CH}_3$ ), 58.17

(SiOCH<sub>2</sub>CH<sub>3</sub>), 58.15 (disubst, SiOCH<sub>2</sub>CH<sub>3</sub>). K+IDS MS (m/e) 693 (M +39, 100%). A small amount of disubstituted product, (CH<sub>2</sub>=CH) ((C<sub>2</sub>H<sub>5</sub>O)<sub>3</sub>SiCH<sub>2</sub>CH<sub>2</sub>)<sub>2</sub>C<sub>6</sub>H<sub>9</sub>, was also observed 529 (M+39, 26%).

5

EXAMPLE 31

## Formation of Very Low Surface Energy Glass

Approximately 0.35 g (0.35 mmol) Star 21 (C<sub>2</sub>H<sub>5</sub>O)<sub>3</sub>Si(CH<sub>2</sub>)<sub>6</sub>(CF<sub>2</sub>)<sub>10</sub>(CH<sub>2</sub>)<sub>6</sub>Si(OC<sub>2</sub>H<sub>5</sub>)<sub>3</sub>, was dissolved in 1.00 g reagent grade tetrahydrofuran to give a  
10 homogeneous solution. Approximately 0.25 g (5.43 mmol) of 96% formic acid was added dropwise to this solution over a period of one minute with moderate stirring. The resultant solution remained clear and formed a clear yellow gel on standing overnight. The yellow color is  
15 believed due to residual platinum catalyst from the synthesis of the Star. The gel was dried over the course of several days at room temperature into a clear yellow glassy disk weighing .278 g. No evidence for open porosity in the glass was obtained when it was  
20 submerged in fluids which wet the surface well.

The surface energy of the glass was assessed by measuring contact angles for several different liquids via the sessile drop method (A. W. Anderson, Physical Chemistry of Surfaces, 4th ed., Wiley-Interscience, NY,  
25 1982, pp. 341-342). After these measurements, the sample was exposed to the silylating agent bis-trimethylsilyl acetamide (a 10% solution by weight in acetonitrile) for 15 min. at room temperature so as to convert residual high energy silanol (Si-OH) surface  
30 groups into Si-O-SiMe<sub>3</sub> groups. The contact angle measurements were then repeated. The very high contact angles for water and methylene iodide indicate a highly hydrophobic low energy surface, especially after the silylation reaction. The surface energy was calculated  
35 to be 15.2 mN/meter, substantially less than that for

poly(tetrafluoroethylene) [Teflon®]. The surface energy was calculated according to the equation

$$\cos \theta = -1 + 2(\gamma_l^d \gamma_s^d)^{1/2} / \gamma_l^d$$

5

where the superscript d refers to the dispersive component of the liquid or solid free energies  $\gamma$ , and is listed in Table II. The equation, which is quite accurate in predicting contact angles of both polar and non-polar liquids on polymers, is based on the assumptions that the reversible work of adhesion can be approximated by its dispersive component, and that the solid/vapor free energy is negligible. See B. Sauer, J. Adhesion Sci. Tech., 6, 955 (1992) for details.

15

Table II  
Fluoroglass Sample Surface Energy Data

	<u>As</u> <u>Generated</u>	<u>After</u> <u>Silylation</u>	<u>PTFE*</u>
Contact angle water (advancing)	90°	123°	110°
Contact angle CH <sub>2</sub> I <sub>2</sub>	Not Measured	90°	83°
Contact angle n-hexadecane	0	11°	40°
Surface energy mN/meter	ca. 30	15.2	23.9
*Poly(tetrafluoroethylene)			

WHAT IS CLAIMED IS:

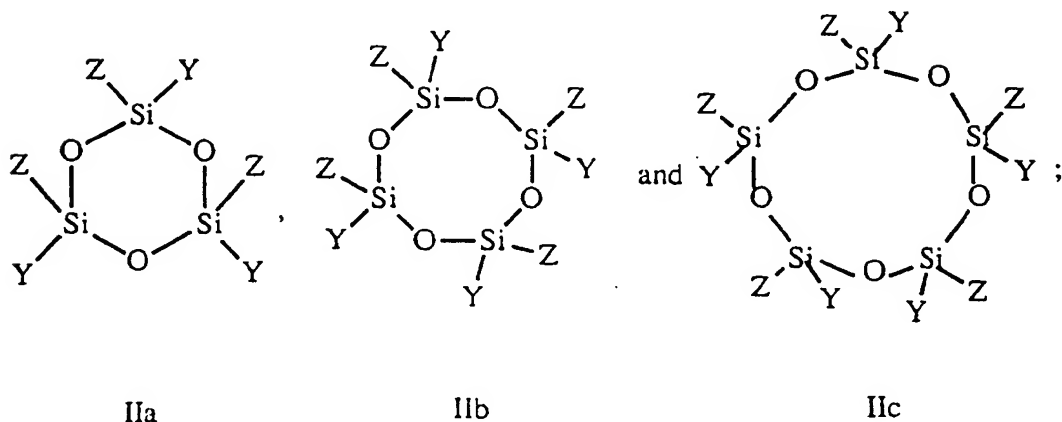
1. An inorganic/organic composition of the idealized empirical formula (II):



wherein

n is an integer greater than or equal to 2; and  
X is at least one flexible organic link selected  
10 from the group consisting of:

- (a)  $\text{R}^1_m\text{SiY}_{4-m}$ ;  
(b) ring structures



15

- (c)  $\text{R}^1_m\text{Si}(\text{OSi}(\text{CH}_3)_2\text{Y})_{4-m}$ ;  
(d)  $\text{R}^1_m\text{Si}(\text{OY})_{4-m}$ ;  
(e)  $\text{CH}_3\text{SiY}_2\text{-O-SiY}_2\text{CH}_3$ ;  
(f)  $\text{Y}(\text{CH}_3)_2\text{Si-C}_6\text{H}_4\text{-Si}(\text{CH}_3)_2\text{Y}$ ;  
20 (g)  $\text{O}[-\text{C}_6\text{H}_4\text{-Si}(\text{CH}_3)_2\text{Y}]_2$ ;  
(h)  $\text{O}[\text{Si}(\text{CH}_3)_2\text{Y}]_2$ ;  
(i)  $\text{Y}(\text{CH}_3)_2\text{SiCH}_2\text{-CH}_2\text{Si}(\text{CH}_3)_2\text{Y}$ ;  
(j)  $\text{Y}(\text{CF}_2)_p\text{Y}$ , provided that when p is 6, Y is  
other than ethylene;  
25 (k)  $\text{Y}_3\text{SiOSiY}_3$ ;  
(l)  $\text{Y}_3\text{Si}(\text{CH}_2)_b\text{SiY}_3$ ;  
(m)  $\text{Y}_3\text{SiC}_6\text{H}_4\text{SiY}_3$ ;

(n) substituted benzene selected from the group consisting of:

- (i)  $C_6H_3(SiZ_{3-a}Y_a)_3$ ;
- (ii)  $C_6H_2(SiZ_{3-a}Y_a)_4$ ;
- 5 (iii)  $C_6H(SiZ_{3-a}Y_a)_5$ ; and
- (iv)  $C_6(SiZ_{3-a}Y_a)_6$ ; and

(o) substituted cyclohexane selected from the group consisting of:

- (i) 1,2- $C_6H_{10}(Y)_2$ ; 1,3- $C_6H_{10}(Y)_2$ ;
- 10 1,4- $C_6H_{10}(Y)_2$ ;
- (ii) 1,2,4- $C_6H_9(Y)_3$ ; 1,2,3- $C_6H_9(Y)_3$ ;
- 1,3,5- $C_6H_9(Y)_3$ ;
- (iii) 1,2,3,4- $C_6H_8(Y)_4$ ; 1,2,4,5- $C_6H_8(Y)_4$ ;
- 1,2,3,5- $C_6H_8(Y)_4$ ;
- 15 (iv) 1,2,3,4,5- $C_6H_7(Y)_5$ ; and
- (v)  $C_6H_6(Y)_6$ ;

wherein:

Z is an alkyl group of 1 to 4 carbon atoms,  
3,3,3-trifluoropropyl, aralkyl, or aryl;

20 Y is  $(CR^2R^3)_kCR^4R^5CR^6R^7(CR^8R^9)_h-$ ;

$R^1$  is alkyl of 1 to about 8 carbon atoms or aryl;

$R^2$  to  $R^9$  are each independently hydrogen, alkyl of  
1 to about 8 carbon atoms or aryl, provided that  
at least one of  $R^4$  to  $R^7$  is hydrogen;

25 m is 0, 1 or 2;

k and h are each independently an integer from 0 to  
10, provided that at least one of k or h is  
zero;

a is 1, 2 or 3;

30 p is an even integer from 4 to 10; and

b is an integer from 1 to 10.

2. A compound of the formula (I):



wherein:

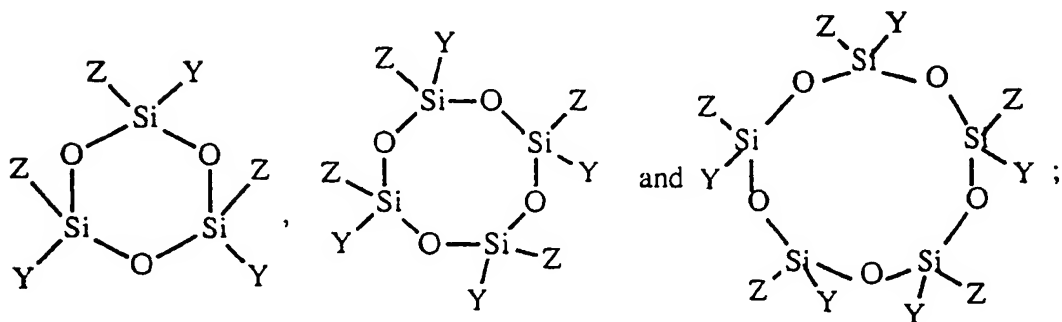
Q is alkoxy of 1 to about 8 carbon atoms, acyloxy of 1 to about 8 carbon atoms, or halogen;

n is an integer greater than or equal to 2; and

5 X is at least one flexible organic link selected from the group consisting of:

(a)  $R^1_mSiY_{4-m}$ ;

(b) ring structures



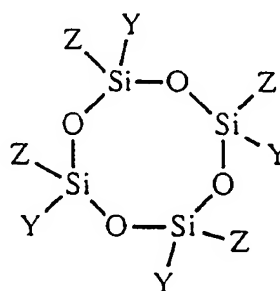
Ia

Ib

Ic

provided that when X is

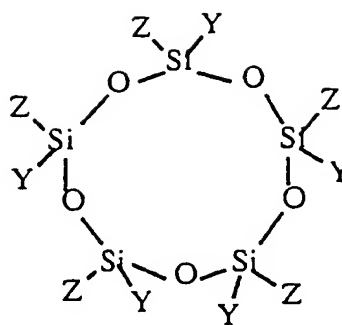
10



Ib

Z is other than methyl and Y is other than ethylene or propylene; and  
15 when X is

53



Ic

Z is other than methyl and Y is other than ethylene or propylene;

5

(c)  $R^1_m \text{Si} (\text{OSi} (\text{CH}_3)_2 \text{Y})_{4-m}$ ;

(d)  $R^1_m \text{Si} (\text{OY})_{4-m}$ ;

(e)  $\text{CH}_3 \text{SiY}_2 - \text{O} - \text{SiY}_2 \text{CH}_3$ ;

(f)  $\text{Y} (\text{CH}_3)_2 \text{Si} - \text{C}_6\text{H}_4 - \text{Si} (\text{CH}_3)_2 \text{Y}$ ;

10

provided that in the definition of Y as defined below either h or k is greater than zero when Q is ethoxy;

(g)  $\text{O} [-\text{C}_6\text{H}_4 - \text{Si} (\text{CH}_3)_2 \text{Y}]_2$ ;

(h)  $\text{O} [\text{Si} (\text{CH}_3)_2 \text{Y}]_2$ ;

15

provided that in the definition of Y as defined below either h or k is greater than zero when Q is ethoxy;

(i)  $\text{Y} (\text{CH}_3)_2 \text{SiCH}_2 - \text{CH}_2 \text{Si} (\text{CH}_3)_2 \text{Y}$ ;

(j)  $\text{Y} (\text{CF}_2)_p \text{Y}$ , provided that Y is other than ethylene;

20

(k)  $\text{Y}_3 \text{SiOSiY}_3$ ;

(l)  $\text{Y}_3 \text{Si} (\text{CH}_2)_b \text{SiY}_3$ ;

(m)  $\text{Y}_3 \text{SiC}_6\text{H}_4 \text{SiY}_3$ ;

(n) substituted benzene selected from the group consisting of:

25

(i)  $\text{C}_6\text{H}_3 (\text{SiZ}_{3-a}\text{Y}_a)_3$ ;

(ii)  $\text{C}_6\text{H}_2 (\text{SiZ}_{3-a}\text{Y}_a)_4$ ;

(iii)  $\text{C}_6\text{H} (\text{SiZ}_{3-a}\text{Y}_a)_5$ ; and

(iv)  $\text{C}_6 (\text{SiZ}_{3-a}\text{Y}_a)_6$ ; and

(o) substituted cyclohexane selected from the group consisting of:

- (i) 1,2-C<sub>6</sub>H<sub>10</sub>(Y)<sub>2</sub>; 1,3-C<sub>6</sub>H<sub>10</sub>(Y)<sub>2</sub>;  
1,4-C<sub>6</sub>H<sub>10</sub>(Y)<sub>2</sub>;  
5 (ii) 1,2,4-C<sub>6</sub>H<sub>9</sub>(Y)<sub>3</sub>; 1,2,3-C<sub>6</sub>H<sub>9</sub>(Y)<sub>3</sub>;  
1,3,5-C<sub>6</sub>H<sub>9</sub>(Y)<sub>3</sub>;  
(iii) 1,2,3,4-C<sub>6</sub>H<sub>8</sub>(Y)<sub>4</sub>; 1,2,4,5-C<sub>6</sub>H<sub>8</sub>(Y)<sub>4</sub>;  
1,2,3,5-C<sub>6</sub>H<sub>8</sub>(Y)<sub>4</sub>;  
(iv) 1,2,3,4,5-C<sub>6</sub>H<sub>7</sub>(Y)<sub>5</sub>; and  
10 (v) C<sub>6</sub>H<sub>6</sub>(Y)<sub>6</sub>;

wherein:

Z is an alkyl group of 1 to 4 carbon atoms,  
3,3,3-trifluoropropyl, aralkyl or aryl;

Y is (CR<sup>2</sup>R<sup>3</sup>)<sub>k</sub>CR<sup>4</sup>R<sup>5</sup>CR<sup>6</sup>R<sup>7</sup>(CR<sup>8</sup>R<sup>9</sup>)<sub>h</sub>-;

15 R<sup>1</sup> is alkyl of 1 to about 8 carbon atoms or  
aryl;

R<sup>2</sup> to R<sup>9</sup> are each independently hydrogen, alkyl  
of 1 to about 8 carbon atoms or aryl,  
provided that at least one of R<sup>4</sup> to R<sup>7</sup> is  
20 hydrogen;

m is 0, 1 or 2;

k and h are each independently an integer from  
0 to 10, provided that at least one of k or  
h is zero;

25 a is 1, 2 or 3;

p is an even integer from 4 to 10; and

b is an integer from 1 to 10.

3. The compound of Claim 2 selected from the  
group consisting of:

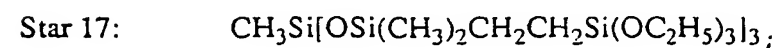
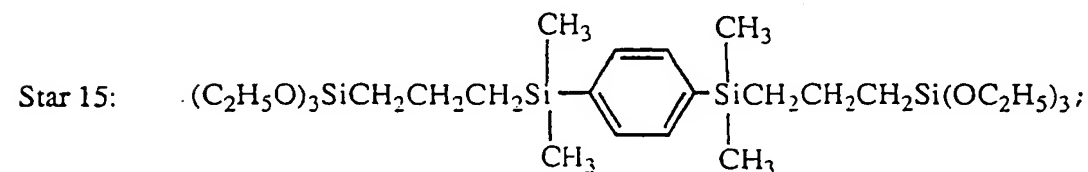
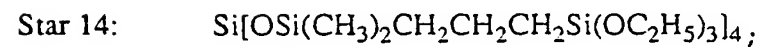
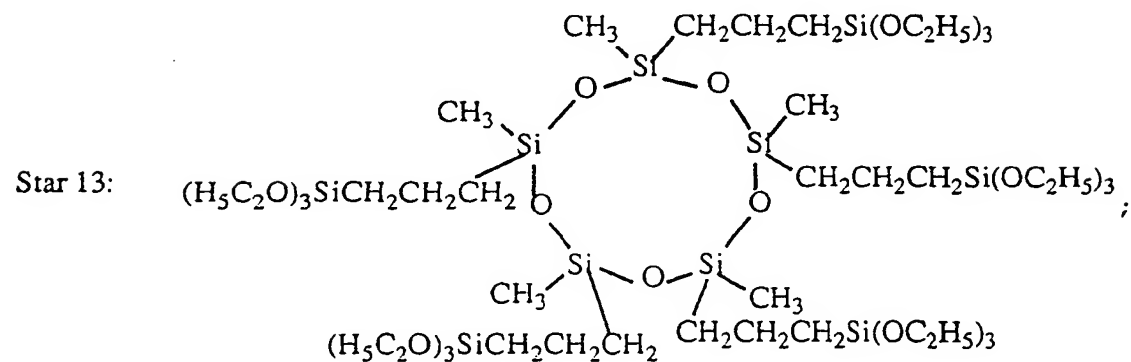
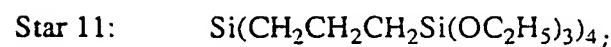
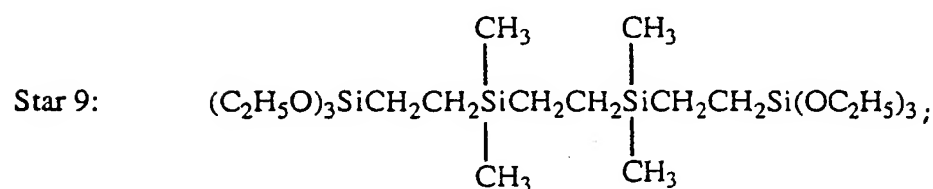
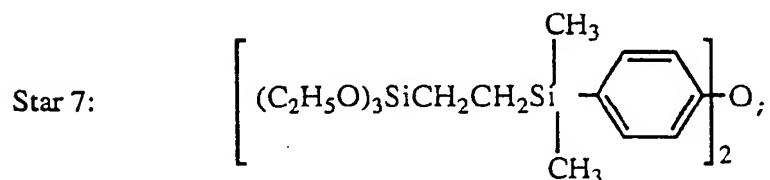
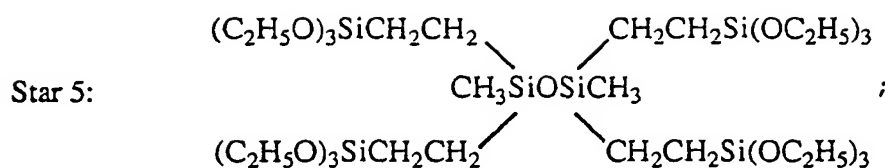
30

Star 1: Si(CH<sub>2</sub>CH<sub>2</sub>Si(OC<sub>2</sub>H<sub>5</sub>)<sub>3</sub>)<sub>4</sub>;

Star 3: Si[OSi(CH<sub>3</sub>)<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Si(OC<sub>2</sub>H<sub>5</sub>)<sub>3</sub>]<sub>4</sub>;

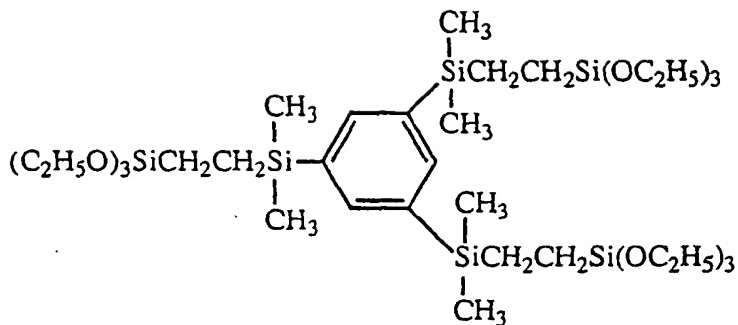
Star 4: Si(OCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Si(OC<sub>2</sub>H<sub>5</sub>)<sub>3</sub>)<sub>4</sub>;



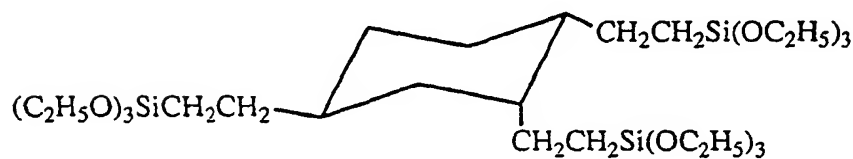


Star 18:  $\text{Si}[\text{CH}_2\text{CH}_2\text{Si}(\text{OCH}_3)_3]_4;$

Star 19:

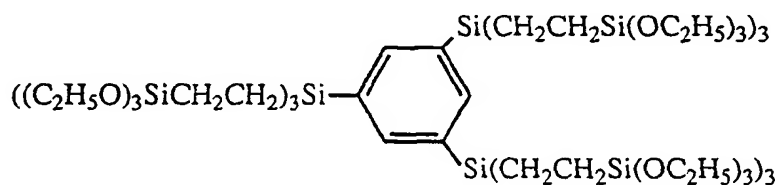


Star 20:

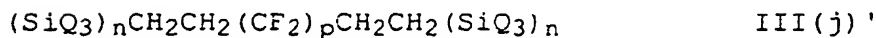


Star 21:  $(\text{C}_2\text{H}_5\text{O})_3\text{Si}(\text{CH}_2)_6(\text{CF}_2)_{10}(\text{CH}_2)_6\text{Si}(\text{OC}_2\text{H}_5)_3;$  and

Star 22:



4. A compound of the formula III(j)':



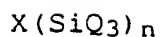
5 wherein:

Q is alkoxy of 1 to about 8 carbon atoms, acyloxy of 1 to about 8 carbon atoms, or halogen;

n is an integer greater than or equal to 2; and

p is an even integer from 4 to 10.

10 5. A process for the preparation of the compound of formula (I):



(I)

wherein:

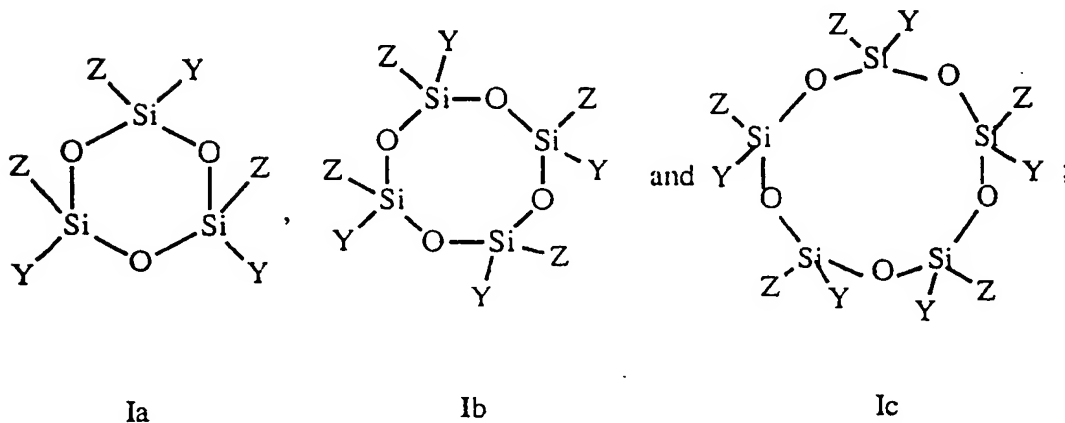
Q is alkoxy of 1 to about 8 carbon atoms, acyloxy  
 5 of 1 to about 8 carbon atoms, or halogen;

n is an integer greater than or equal to 2; and

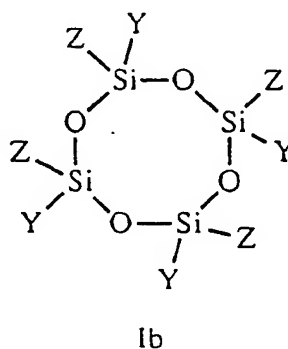
X is at least one flexible organic link selected  
 from the group consisting of:

(a)  $\text{R}^1_m\text{SiY}_{4-m}$ ;

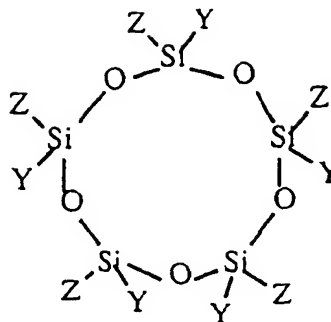
10 (b) ring structures



provided that when X is



Z is other than methyl and Y is other than  
 ethylene or propylene; and  
 when X is



Ic

Z is other than methyl and Y is other than ethylene or propylene;

(c)  $R^1_m \text{Si}(\text{OSi}(\text{CH}_3)_2\text{Y})_{4-m}$ ;

(d)  $R^1_m \text{Si}(\text{OY})_{4-m}$ ;

5

(e)  $\text{CH}_3\text{SiY}_2\text{-O-SiY}_2\text{CH}_3$ ;

(f)  $\text{Y}(\text{CH}_3)_2\text{Si-C}_6\text{H}_4\text{-Si}(\text{CH}_3)_2\text{Y}$ ;

provided that either h or k is greater than zero when Q is ethoxy;

(g)  $\text{O}[-\text{C}_6\text{H}_4\text{-Si}(\text{CH}_3)_2\text{Y}]_2$ ;

10

(h)  $\text{O}[\text{Si}(\text{CH}_3)_2(\text{Y})]_2$ ;

provided that either h or k is greater than zero when Q is ethoxy;

(i)  $\text{Y}(\text{CH}_3)_2\text{SiCH}_2\text{-CH}_2\text{Si}(\text{CH}_3)_2\text{Y}$ ;

(j)  $\text{Y}(\text{CF}_2)_p\text{Y}$ , provided that Y is other than ethylene;

15

(k)  $\text{Y}_3\text{SiOSiY}_3$ ;

(l)  $\text{Y}_3\text{Si}(\text{CH}_2)_b\text{SiY}_3$ ;

(m)  $\text{Y}_3\text{SiC}_6\text{H}_4\text{SiY}_3$ ;

(n) substituted benzene selected from the group consisting of:

20

(i)  $\text{C}_6\text{H}_3(\text{SiZ}_{3-a}\text{Y}_a)_3$ ;

(ii)  $\text{C}_6\text{H}_2(\text{SiZ}_{3-a}\text{Y}_a)_4$ ;

(iii)  $\text{C}_6\text{H}(\text{SiZ}_{3-a}\text{Y}_a)_5$ ; and

(iv)  $\text{C}_6(\text{SiZ}_{3-a}\text{Y}_a)_6$ ; and

(o) substituted cyclohexane selected from the group consisting of:

- (i) 1,2-C<sub>6</sub>H<sub>10</sub>(Y)<sub>2</sub>; 1,3-C<sub>6</sub>H<sub>10</sub>(Y)<sub>2</sub>;  
1,4-C<sub>6</sub>H<sub>10</sub>(Y)<sub>2</sub>
- 5 (ii) 1,2,4-C<sub>6</sub>H<sub>9</sub>(Y)<sub>3</sub>; 1,2,3-C<sub>6</sub>H<sub>9</sub>(Y)<sub>3</sub>;  
1,3,5-C<sub>6</sub>H<sub>9</sub>(Y)<sub>3</sub>;
- (iii) 1,2,3,4-C<sub>6</sub>H<sub>8</sub>(Y)<sub>4</sub>; 1,2,4,5-C<sub>6</sub>H<sub>8</sub>(Y)<sub>4</sub>;  
1,2,3,5-C<sub>6</sub>H<sub>8</sub>(Y)<sub>4</sub>;
- (iv) 1,2,3,4,5-C<sub>6</sub>H<sub>7</sub>(Y)<sub>5</sub>; and
- 10 (v) C<sub>6</sub>H<sub>6</sub>(Y)<sub>6</sub>;

wherein:

Z is an alkyl group of 1 to 4 carbon atoms,  
3,3,3-trifluoropropyl, aralkyl, or aryl;

Y is (CR<sup>2</sup>R<sup>3</sup>)<sub>k</sub>CR<sup>4</sup>R<sup>5</sup>CR<sup>6</sup>R<sup>7</sup>-;

15 R<sup>1</sup> is alkyl of 1 to about 8 carbon atoms or  
aryl;

R<sup>2</sup> to R<sup>9</sup> are each independently hydrogen, alkyl  
of 1 to about 8 carbon atoms or aryl,  
provided that at least one of R<sup>4</sup> to R<sup>7</sup> is  
20 hydrogen;

m is 0, 1 or 2;

k and h are each independently an integer from  
0 to 10, provided that at least one of k or  
h is zero;

25 a is 1, 2 or 3;

p is an even integer from 4 to 10; and

b is an integer from 1 to 10;

comprising reacting a compound containing an Si-H group  
with a compound containing an olefinic or alkynyl bond  
30 in the presence of a transition metal catalyst or free  
radical initiator.

6. A process for the preparation of the  
composition of formula (II)

35 
$$X(SiO_{1.5})_n \quad (II)$$

wherein

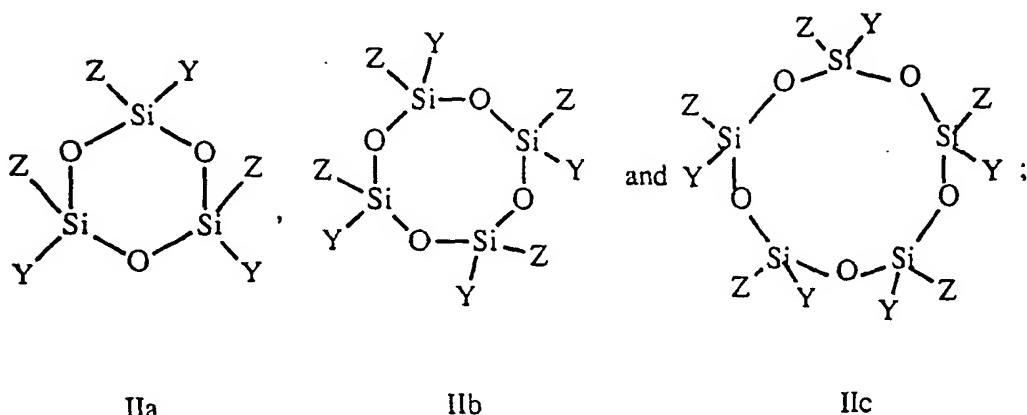
$n$  is an integer greater than or equal to 2; and

$X$  is at least one flexible organic link selected

5 from the group consisting of:

(a)  $R^1_m Si Y_{4-m}$ ;

(b) ring structures



10

(c)  $R^1_m Si (OSi (CH_3)_2 Y)_{4-m}$ ;

(d)  $R^1_m Si (OY)_{4-m}$ ;

(e)  $CH_3 Si Y_2 - O - Si Y_2 CH_3$ ;

(f)  $Y (CH_3)_2 Si - C_6H_4 - Si (CH_3)_2 Y$ ;

15

(g)  $O [-C_6H_4 - Si (CH_3)_2 Y]_2$ ;

(h)  $O [Si (CH_3)_2 Y]_2$ ;

(i)  $Y (CH_3)_2 Si CH_2 - CH_2 Si (CH_3)_2 Y$ ;

(j)  $Y (CF_2)_p Y$ , provided that when  $p$  is 6,  $Y$  is other than ethylene;

20

(k)  $Y_3 Si O Si Y_3$ ;

(l)  $Y_3 Si (CH_2)_b Si Y_3$ ;

(m)  $Y_3 Si C_6H_4 Si Y_3$ ;

(n) substituted benzene selected from the group consisting of:

25

(i)  $C_6H_3 (Si Z_{3-a} Y_a)_3$ ;

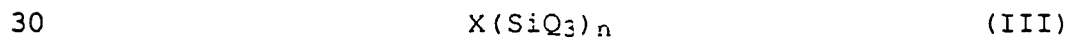
(ii)  $C_6H_2 (Si Z_{3-a} Y_a)_4$ ;

(iii)  $C_6H (Si Z_{3-a} Y_a)_5$ ; and

- (iv)  $C_6(SiZ_{3-a}Y_a)_6$ ; and
- (o) substituted cyclohexane selected from the group consisting of:
- (i) 1,2- $C_6H_{10}(Y)_2$ ; 1,3- $C_6H_{10}(Y)_2$ ;  
 1,4- $C_6H_{10}(Y)_2$ ;
  - (ii) 1,2,4- $C_6H_9(Y)_3$ ; 1,2,3- $C_6H_9(Y)_3$ ;  
 1,3,5- $C_6H_9(Y)_3$ ;
  - (iii) 1,2,3,4- $C_6H_8(Y)_4$ ; 1,2,4,5- $C_6H_8(Y)_4$ ;  
 1,2,3,5- $C_6H_8(Y)_4$ ;
  - (iv) 1,2,3,4,5- $C_6H_7(Y)_5$ ; and
  - (v)  $C_6H_6(Y)_6$ ;
- wherein:
- Z is an alkyl group of 1 to 4 carbon atoms,  
 3,3,3-trifluoropropyl, aralkyl, or aryl;
- Y is  $(CR^2R^3)_kCR^4R^5CR^6R^7(CR^8R^9)_h-$ ;
- $R^1$  is alkyl of 1 to about 8 carbon atoms or aryl;  
 $R^2$  to  $R^9$  are each independently hydrogen, alkyl of  
 1 to about 8 carbon atoms or aryl, provided that  
 at least one of  $R^4$  to  $R^7$  is hydrogen;
- m is 0, 1 or 2;
- k and h are each independently an integer from 0 to  
 10, provided that at least one of k or h is  
 zero;
- a is 1, 2 or 3;
- p is an even integer from 4 to 10; and
- b is an integer from 1 to 10

comprising:

(A) mixing at least one compound of formula (III):



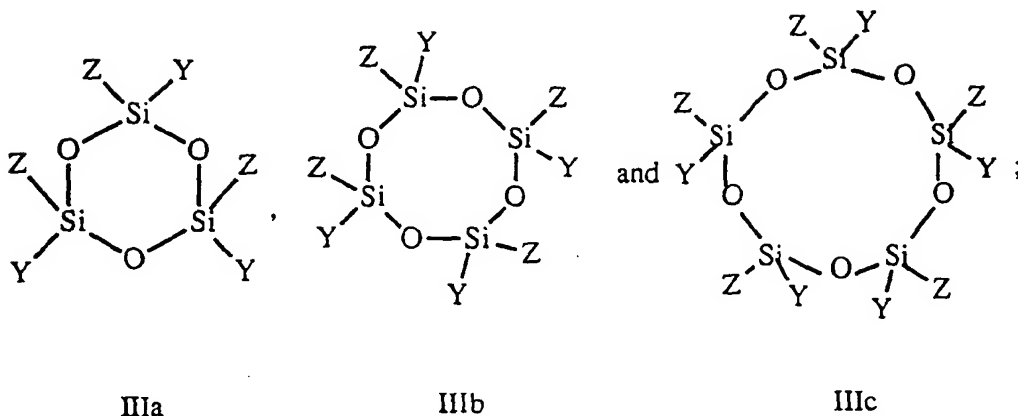
wherein:

Q is alkoxy of 1 to about 8 carbon atoms, acyloxy  
 of 1 to about 8 carbon atoms, or halogen;

n is an integer greater than or equal to 2; and

X is at least one flexible organic link selected from the group consisting of:

- (a)  $R^1_m Si Y_{4-m}$ ;
- (b) ring structures



- 5 (c)  $R^1_m Si (OSi (CH_3)_2 Y)_{4-m}$ ;
- (d)  $R^1_m Si (OY)_{4-m}$ ;
- (e)  $CH_3 Si Y_2 - O - Si Y_2 CH_3$ ;
- (f)  $Y (CH_3)_2 Si - C_6H_4 - Si (CH_3)_2 Y$ ;
- (g)  $O [-C_6H_4 - Si (CH_3)_2 Y]_2$ ;
- 10 (h)  $O [Si (CH_3)_2 Y]_2$ ;
- (i)  $Y (CH_3)_2 Si CH_2 - CH_2 Si (CH_3)_2 Y$ ;
- (j)  $Y (CF_2)_p Y$ , provided that when p is 6, Y is other than ethylene;
- (k)  $Y_3 Si O Si Y_3$ ;
- 15 (l)  $Y_3 Si (CH_2)_5 Si Y_3$ ;
- (m)  $Y_3 Si C_6H_4 Si Y_3$ ;
- (n) substituted benzene selected from the group consisting of:
  - (i)  $C_6H_3 (Si Z_{3-a} Y_a)_3$ ;
  - 20 (ii)  $C_6H_2 (Si Z_{3-a} Y_a)_4$ ;
  - (iii)  $C_6H (Si Z_{3-a} Y_a)_5$ ; and
  - (iv)  $C_6 (Si Z_{3-a} Y_a)_6$ ; and



(o) substituted cyclohexane selected from the group consisting of:

- (i) 1,2-C<sub>6</sub>H<sub>10</sub>(Y)<sub>2</sub>; 1,3-C<sub>6</sub>H<sub>10</sub>(Y)<sub>2</sub>;  
1,4-C<sub>6</sub>H<sub>10</sub>(Y)<sub>2</sub>  
5 (ii) 1,2,4-C<sub>6</sub>H<sub>9</sub>(Y)<sub>3</sub>; 1,2,3-C<sub>6</sub>H<sub>9</sub>(Y)<sub>3</sub>;  
1,3,5-C<sub>6</sub>H<sub>9</sub>(Y)<sub>3</sub>;  
(iii) 1,2,3,4-C<sub>6</sub>H<sub>8</sub>(Y)<sub>4</sub>; 1,2,4,5-C<sub>6</sub>H<sub>8</sub>(Y)<sub>4</sub>;  
1,2,3,5-C<sub>6</sub>H<sub>8</sub>(Y)<sub>4</sub>;  
(iv) 1,2,3,4,5-C<sub>6</sub>H<sub>7</sub>(Y)<sub>5</sub>; and  
10 (v) C<sub>6</sub>H<sub>6</sub>(Y)<sub>6</sub>;

wherein:

Z is an alkyl group of 1 to 4 carbon atoms,  
3,3,3-trifluoropropyl, aralkyl, or aryl;

Y is (CR<sup>2</sup>R<sup>3</sup>)<sub>k</sub>CR<sup>4</sup>R<sup>5</sup>CR<sup>6</sup>R<sup>7</sup>(CR<sup>8</sup>R<sup>9</sup>)<sub>h</sub>-;

- 15 R<sup>1</sup> is alkyl of 1 to about 8 carbon atoms or aryl;  
R<sup>2</sup> to R<sup>9</sup> are each independently hydrogen, alkyl of  
1 to about 8 carbon atoms or aryl, provided that  
at least one of R<sup>4</sup> to R<sup>7</sup> is hydrogen;

m is 0, 1 or 2;

- 20 k and h are each independently an integer from 0 to  
10, provided that at least one of k or h is  
zero;

a is 1, 2 or 3;

p is an even integer from 4 to 10; and

- 25 b is an integer from 1 to 10,

with water in the presence of a solvent and a catalyst,  
or at least one carboxylic acid having a maximum pKa  
value about 4.0 and containing from 0 to 20 mole %  
water;

- 30 (B) maintaining the mixture resulting from step  
(A) at a temperature within the range of about  
0-100°C; and  
(C) isolating the resulting inorganic/organic  
composition of formula (II).

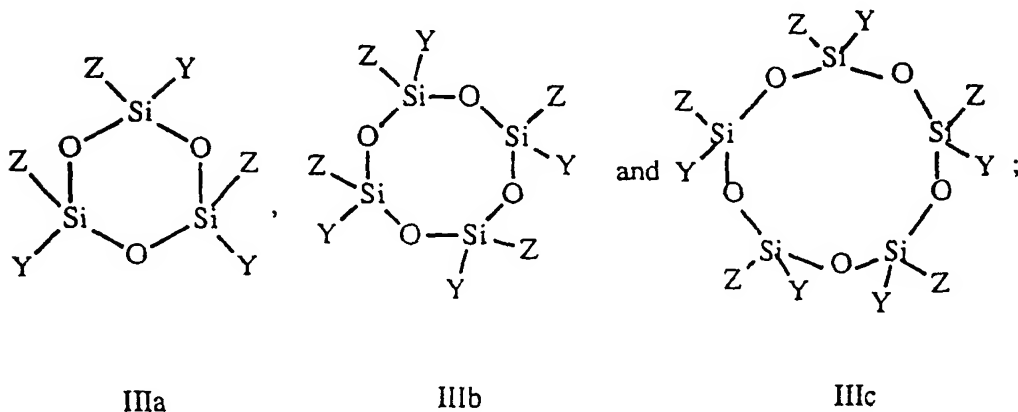
7. A method for modifying a sol-gel glass to generate a sol-gel glass that can tolerate increased drying rates and shows lower brittleness comprising:

- (A) combining a star gel precursor of formula  
5 (III):



wherein:

- 10 Q is alkoxy of 1 to about 8 carbon atoms, acyloxy of 1 to about 8 carbon atoms, or halogen;  
n is an integer greater than or equal to 2; and  
X is at least one flexible organic link selected from the group consisting of:  
15 (a)  $\text{R}^1_m\text{SiY}_{4-m}$ ;  
(b) ring structures



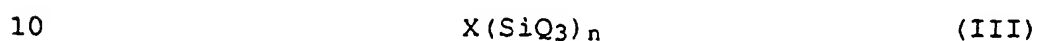
- (c)  $\text{R}^1_m\text{Si}(\text{OSi}(\text{CH}_3)_2\text{Y})_{4-m}$ ;  
(d)  $\text{R}^1_m\text{Si}(\text{OY})_{4-m}$ ;  
(e)  $\text{CH}_3\text{SiY}_2\text{-O-SiY}_2\text{CH}_3$ ;  
20 (f)  $\text{Y}(\text{CH}_3)_2\text{Si-C}_6\text{H}_4\text{-Si}(\text{CH}_3)_2\text{Y}$ ;  
(g)  $\text{O}[-\text{C}_6\text{H}_4\text{-Si}(\text{CH}_3)_2\text{Y}]_2$ ;  
(h)  $\text{O}[\text{Si}(\text{CH}_3)_2\text{Y}]_2$ ;  
(i)  $\text{Y}(\text{CH}_3)_2\text{SiCH}_2\text{-CH}_2\text{Si}(\text{CH}_3)_2\text{Y}$ ;  
(j)  $\text{Y}(\text{CF}_2)_p\text{Y}$ ;

- (k)  $Y_3SiOSiY_3$ ;  
 (l)  $Y_3Si(CH_2)_bSiY_3$ ;  
 (m)  $Y_3SiC_6H_4SiY_3$ ;  
 (n) substituted benzene selected from the group  
 5 consisting of:  
     (i)  $C_6H_3(SiZ_{3-a}Y_a)_3$ ;  
     (ii)  $C_6H_2(SiZ_{3-a}Y_a)_4$ ;  
     (iii)  $C_6H(SiZ_{3-a}Y_a)_5$ ; and  
     (iv)  $C_6(SiZ_{3-a}Y_a)_6$ ; and  
 10 (o) substituted cyclohexane selected from the  
     group consisting of:  
     (i) 1,2- $C_6H_{10}(Y)_2$ ; 1,3- $C_6H_{10}(Y)_2$ ;  
         1,4- $C_6H_{10}(Y)_2$   
     (ii) 1,2,4- $C_6H_9(Y)_3$ ; 1,2,3- $C_6H_9(Y)_3$ ;  
 15 1,3,5- $C_6H_9(Y)_3$ ;  
     (iii) 1,2,3,4- $C_6H_8(Y)_4$ ; 1,2,4,5- $C_6H_8(Y)_4$ ;  
         1,2,3,5- $C_6H_8(Y)_4$ ;  
     (iv) 1,2,3,4,5- $C_6H_7(Y)_5$ ; and  
     (v)  $C_6H_6(Y)_6$ ;  
 20 wherein:  
     Z is an alkyl group of 1 to 4 carbon atoms,  
         3,3,3-trifluoropropyl, aralkyl, or aryl;  
     Y is  $(CR^2R^3)_kCR^4R^5CR^6R^7(CR^8R^9)_h-$ ;  
      $R^1$  is alkyl of 1 to about 8 carbon atoms or  
 25 aryl;  
      $R^2$  to  $R^9$  are each independently hydrogen, alkyl  
         of 1 to about 8 carbon atoms or aryl,  
         provided that at least one of  $R^4$  to  $R^7$  is  
         hydrogen;  
 30 m is 0, 1 or 2;  
     k and h are each independently an integer from  
         0 to 10, provided that least one of k or h  
         is zero;  
     a is 1, 2 or 3;  
 35 p is an even integer from 4 to 10; and

b is an integer from 1 to 10;  
with a metal alkoxide sol-gel precursor;

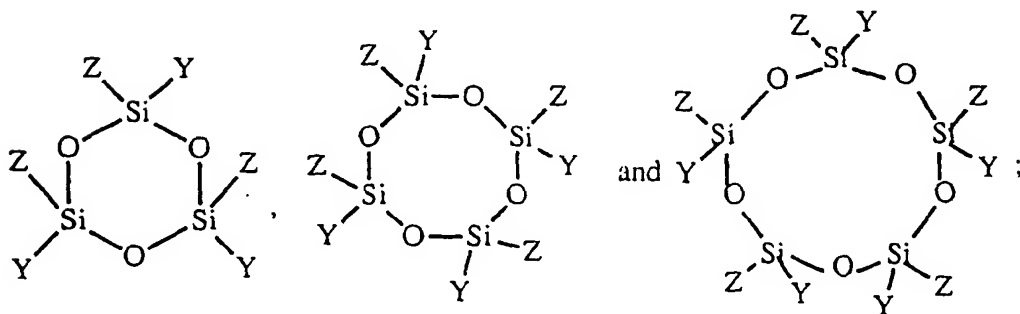
- (B) mixing in water with a solvent and a catalyst, or a carboxylic acid optionally in the presence of a solvent; and  
5 (C) drying.

8. A method for coating a substrate comprising reacting the star gel precursor of formula (III):



wherein:

- Q is alkoxy of 1 to about 8 carbon atoms, acyloxy of 1 to about 8 carbon atoms, or halogen;  
15 n is an integer greater than or equal to 2; and  
X is at least one flexible organic link selected from the group consisting of:  
(a)  $\text{R}^1_m\text{Si}(\text{Y})_{4-m}$ ;  
(b) ring structures



IIIa

IIIb

IIIc

- 20 (c)  $\text{R}^1_m\text{Si}(\text{OSi}(\text{CH}_3)_2\text{Y})_{4-m}$ ;  
(d)  $\text{R}^1_m\text{Si}(\text{OY})_{4-m}$ ;  
(e)  $\text{CH}_3\text{SiY}_2\text{-O-SiY}_2\text{CH}_3$ ;  
(f)  $\text{Y}(\text{CH}_3)_2\text{Si-C}_6\text{H}_4\text{-Si}(\text{CH}_3)_2\text{Y}$ ;  
(g)  $\text{O}[-\text{C}_6\text{H}_4\text{-Si}(\text{CH}_3)_2\text{Y}]_2$ ;

- (h)  $O[Si(CH_3)_2Y]_2$ ;  
 (i)  $Y(CH_3)_2SiCH_2-CH_2Si(CH_3)_2Y$ ;  
 (j)  $Y(CF_2)_pY$ ;  
 (k)  $Y_3SiOSiY_3$ ;  
 5 (l)  $Y_3Si(CH_2)_bSiY_3$ ;  
 (m)  $Y_3SiC_6H_4SiY_3$ ;  
 (n) substituted benzene selected from the group consisting of:  
 (i)  $C_6H_3(SiZ_{3-a}Y_a)_3$ ;  
 10 (ii)  $C_6H_2(SiZ_{3-a}Y_a)_4$ ;  
 (iii)  $C_6H(SiZ_{3-a}Y_a)_5$ ; and  
 (iv)  $C_6(SiZ_{3-a}Y_a)_6$ ; and  
 (o) substituted cyclohexane selected from the group consisting of:  
 15 (i) 1,2- $C_6H_{10}(Y)_2$ ; 1,3- $C_6H_{10}(Y)_2$ ;  
 1,4- $C_6H_{10}(Y)_2$ ;  
 (ii) 1,2,4- $C_6H_9(Y)_3$ ; 1,2,3- $C_6H_9(Y)_3$ ;  
 1,3,5- $C_6H_9(Y)_3$ ;  
 (iii) 1,2,3,4- $C_6H_8(Y)_4$ ; 1,2,4,5- $C_6H_8(Y)_4$ ;  
 20 1,2,3,5- $C_6H_8(Y)_4$ ;  
 (iv) 1,2,3,4,5- $C_6H_7(Y)_5$ ; and  
 (v)  $C_6H_6(Y)_6$ ;

wherein:

- Z is an alkyl group of 1 to 4 carbon atoms,  
 25 3,3,3-trifluoropropyl, aralkyl or aryl;  
 Y is  $(CR^2R^3)_kCR^4R^5CR^6R^7(CR^8R^9)_h-$ ;  
 $R^1$  is alkyl of 1 to about 8 carbon atoms or aryl;  
 $R^2$  to  $R^9$  are each independently hydrogen, alkyl  
 30 of 1 to about 8 carbon atoms or aryl,  
 provided that at least one of  $R^4$  to  $R^7$  is hydrogen;  
 m is 0, 1 or 2;

k and h are each independently an integer from 0 to 10, provided that at least one of k or h is zero; and

a is 1, 2 or 3;

5 p is an even integer from 4 to 10; and

b is an integer from 1 to 10;

with water in the presence of a solvent and a catalyst, or a strong carboxylic acid, optionally in the presence of a solvent; dipping the substrate in the resulting mixture; removing the coated substrate from the mixture and drying the coating to generate a substrate coated with a composition of formula II



15

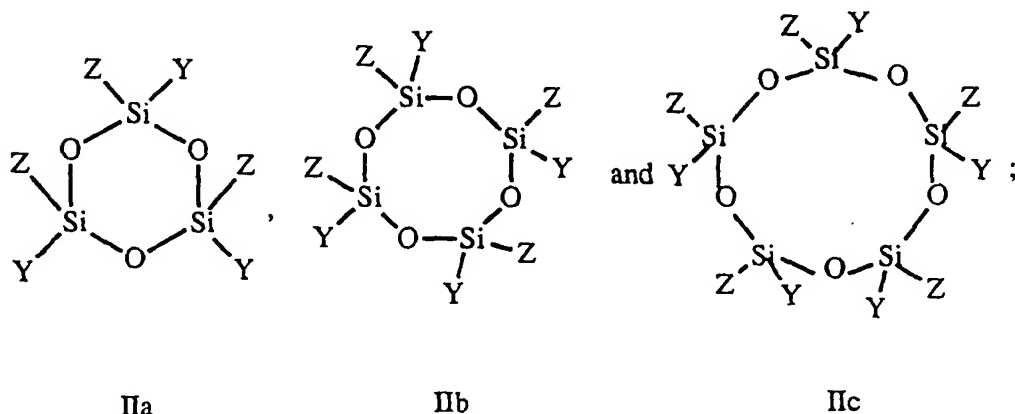
wherein

n is an integer greater than or equal to 2; and X is at least one flexible organic link selected from the group consisting of:

20

(a)  $\text{R}^1_m\text{SiY}_{4-m}$ ;

(b) ring structures



25

(c)  $\text{R}^1_m\text{Si}(\text{OSi}(\text{CH}_3)_2\text{Y})_{4-m}$ ;

(d)  $\text{R}^1_m\text{Si}(\text{OY})_{4-m}$ ;

(e)  $\text{CH}_3\text{SiY}_2\text{-O-SiY}_2\text{CH}_3$ ;

- (f)  $Y(CH_3)_2Si-C_6H_4-Si(CH_3)_2Y$ ;  
 (g)  $O[-C_6H_4-Si(CH_3)_2Y]_2$ ;  
 (h)  $O[Si(CH_3)_2Y]_2$ ;  
 (i)  $Y(CH_3)_2SiCH_2-CH_2Si(CH_3)_2Y$ ;  
 5 (j)  $Y(CF_2)_pY$ , provided that when p is 6, Y is other than ethylene;  
 (k)  $Y_3SiOSiY_3$ ;  
 (l)  $Y_3Si(CH_2)_bSiY_3$ ;  
 (m)  $Y_3SiC_6H_4SiY_3$ ;  
 10 (n) substituted benzene selected from the group consisting of:  
     (i)  $C_6H_3(SiZ_{3-a}Y_a)_3$ ;  
     (ii)  $C_6H_2(SiZ_{3-a}Y_a)_4$ ;  
     (iii)  $C_6H(SiZ_{3-a}Y_a)_5$ ; and  
 15 (iv)  $C_6(SiZ_{3-a}Y_a)_6$ ; and  
 (o) substituted cyclohexane selected from the group consisting of:  
     (i) 1,2- $C_6H_{10}(Y)_2$ ; 1,3- $C_6H_{10}(Y)_2$ ;  
         1,4- $C_6H_{10}(Y)_2$ ;  
 20 (ii) 1,2,4- $C_6H_9(Y)_3$ ; 1,2,3- $C_6H_9(Y)_3$ ;  
         1,3,5- $C_6H_9(Y)_3$ ;  
     (iii) 1,2,3,4- $C_6H_8(Y)_4$ ; 1,2,4,5- $C_6H_8(Y)_4$ ;  
         1,2,3,5- $C_6H_8(Y)_4$ ;  
     (iv) 1,2,3,4,5- $C_6H_7(Y)_5$ ; and  
 25 (v)  $C_6H_6(Y)_6$ ;

wherein:

Z is an alkyl group of 1 to 4 carbon atoms,

3,3,3-trifluoropropyl, aralkyl, or aryl;

Y is  $(CR^2R^3)_kCR^4R^5CR^6R^7(CR^8R^9)_h-$ ;

30  $R^1$  is alkyl of 1 to about 8 carbon atoms or aryl;

$R^2$  to  $R^9$  are each independently hydrogen, alkyl of 1 to about 8 carbon atoms or aryl, provided that at least one of  $R^4$  to  $R^7$  is hydrogen;

m is 0, 1 or 2;

70

k and h are each independently an integer from 0 to 10, provided that at least one of k or h is zero;

a is 1, 2 or 3;

5

p is an even integer from 4 to 10; and

b is an integer from 1 to 10.



## INTERNATIONAL SEARCH REPORT

International Application No

PCT/US 93/08685

## A. CLASSIFICATION OF SUBJECT MATTER

IPC 5 C07F7/18 C07F7/12 C07F7/21 C08G77/50 C08G77/52  
C09D183/14

According to International Patent Classification (IPC) or to both national classification and IPC

## B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 5 C07F C08G C09D

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

## C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	US,A,4 461 867 (SURPRENANT, R.) 24 July 1984	2,3
A	see the whole document	
	see the whole document	1,5-8
	---	
X	US,A,5 145 907 (KALINOWSKI, R.E. ET AL.) 8 September 1992	2,3
	see the whole document	
A	see the whole document	1,5-8
	---	
	-/--	

☒ Further documents are listed in the continuation of box C.☒ Patent family members are listed in annex.

## \* Special categories of cited documents :

- \*A\* document defining the general state of the art which is not considered to be of particular relevance
- \*E\* earlier document but published on or after the international filing date
- \*L\* document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)
- \*O\* document referring to an oral disclosure, use, exhibition or other means
- \*P\* document published prior to the international filing date but later than the priority date claimed

- \*T\* later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
- \*X\* document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
- \*Y\* document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.
- \*&\* document member of the same patent family

Date of the actual completion of the international search

11 November 1993

Date of mailing of the international search report

30. 11. 93

Name and mailing address of the ISA

European Patent Office, P.B. 5818 Patentlaan 2  
NL - 2280 HV Rijswijk  
Tel. (+ 31-70) 340-2040, Tx. 31 651 epo nl,  
Fax (+ 31-70) 340-3016

Authorized officer

RINKEL, L

## INTERNATIONAL SEARCH REPORT

Int: mal Application No

PCT/US 93/08685

## C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	CHEMICAL ABSTRACTS, vol. 115, no. 20, 18 November 1991, Columbus, Ohio, US; abstract no. 210341f, OSUGI, K. ET AL. 'SILOXANE CROSSLINKER-CONTAINING COATINGS' page 127 ;	2,3
A	& JP,A,0 343 454 (NIPPON PAINT CO., LTD.; SHIN-ETSU CHEMICAL INDUSTRY CO., LTD.) 25 February 1991	1,5-8
X	--- US,A,4 689 085 (PLUEDDEMAN, E.P.) 25 August 1987 see the whole document	2,3
X	--- EP,A,0 496 597 (SHIN-ETSU CHEMICAL CO., LTD.) 29 July 1992 see the whole document	2,3
X	--- EP,A,0 183 533 (TORAY SILICONE CO., LTD.) 4 June 1986 see the whole document	2,3
X	--- CHEMICAL ABSTRACTS, vol. 108, no. 10, 7 March 1988, Columbus, Ohio, US; abstract no. 76116m, ONA, I. ET AL. 'ALKOXYSILYLALKYL/POLYOXYALKYLENE MODIFIED ORGANOPOLYSILOXANES' page 17 ; see abstract & JP,A,62 225 533 (TORAY SILICONE CO., LTD.) 3 October 1987	2,3
X	--- CHEMICAL ABSTRACTS, vol. 110, no. 8, 20 February 1989, Columbus, Ohio, US; abstract no. 64253g, LEWIS, L.N. ET AL. 'PREPARATION AND STRUCTURE OF PLATINUM GROUP METAL COLLOIDS: WITHOUT SOLVENTS' page 436 ; see abstract & CHEM. MATER. vol. 1, no. 1, 1989 pages 106 - 114	2,3
X	--- EP,A,0 409 272 (DOW CORNING CORPORATION) 23 January 1991 see the whole document	2,3
	--- -/--	

1

## INTERNATIONAL SEARCH REPORT

International Application No

PCT/US 93/08685

## C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	CHEMICAL ABSTRACTS, vol. 110, no. 20, 15 May 1989, Columbus, Ohio, US; abstract no. 174875u, GORSHKOV, A.V. ET AL. 'EFFECT OF OLIGOALKOXYSILOXANES ON THE PROPERTIES OF SKTV-M VULCANIZATES' page 81 ; see abstract & KAUCH. REZINA no. 9 , 1988 pages 11 - 15 ----	2,3
X	JOURNAL OF ORGANOMETALLIC CHEMISTRY vol. 299 , 1986 BIRKOFER, L. ET AL. 'SILICIUMORGANISCHE VERBINDUNGEN. LXXXVIII. DISILATRANYLALKANE' see the whole document ----	2,3
X	BULLETIN DE LA SOCIETE CHIMIQUE DE FRANCE no. 11 , 1971 pages 4107 - 4116 MICHEL, C. ET AL. 'SYNTHESE, PROPRIETES SPECTROSCOPIQUES IR ET RMI ET VISCOSIMETRIE DE QUELQUES P-PHENYLENE DISILANES SUBSTITUES' see the whole document ----	2,3
A	CHEMICAL ABSTRACTS, vol. 116, no. 6, 10 February 1992, Columbus, Ohio, US; abstract no. 42881g, KAGAWA, H. ET AL. 'EPOXY RESIN COMPOSITIONS FOR SEMICONDUCTOR SEALANTS' page 63 ; see abstract & JP,A,03 192 151 (MATSUSHITA ELECTRIC WORKS, LTD.) 22 August 1991 ----	2,3
X	DE,A,19 66 503 (DOW CORNING CORPORATION) 8 March 1973 see examples 1-22 ----	4
P,X	US,A,5 202 404 (SHIN-ETSU CHEMICAL CO., LTD.) 13 April 1993 see the whole document -----	1-8

**INTERNATIONAL SEARCH REPORT**  
information on patent family members

International Application No  
**PCT/US 93/08685**

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
US-A-4461867	24-07-84	JP-C- 1712095 JP-B- 4000108 JP-A- 59070694 JP-A- 3193792 US-A- 4525400	11-11-92 06-01-92 21-04-84 23-08-91 25-06-85
US-A-5145907	08-09-92	US-A- 4954565 AU-A- 6309190 EP-A- 0420522 JP-A- 3122163	04-09-90 28-03-91 03-04-91 24-05-91
JP-A-0343454		NONE	
US-A-4689085	25-08-87	EP-A- 0255227	03-02-88
EP-A-0496597	29-07-92	JP-A- 4243884	31-08-92
EP-A-0183533	04-06-86	JP-B- 5017927 JP-A- 61127733 CA-A- 1272207 US-A- 4616076	10-03-93 16-06-86 31-07-90 07-10-86
JP-A-62225533	03-10-87	NONE	
EP-A-0409272	23-01-91	AU-B- 622569 AU-A- 5917490 JP-A- 3128969	09-04-92 24-01-91 31-05-91
JP-A-03192151	22-08-91	NONE	
DE-A-1966503	08-03-73	AT-B- 320280 AT-A, B 305309 BE-A- 736871 CH-A- 517822 DE-A- 1939300 FR-A- 2014927 GB-A- 1230931 NL-A- 6911798 SE-B- 357751 US-A- 3542830	10-02-75 15-01-73 02-02-70 15-01-72 19-02-70 24-04-70 05-05-71 04-02-70 09-07-73 24-11-70

information on patent family members

PCT/US 93/08685

Form PCT/ISA/210 (patent family annex) (July 1992)

